

Worksheet- Calculations Involving Specific Heat

1. For $q = m \cdot c \cdot \Delta T$: identify each variables by name & the units associated with it.

q = amount of heat (J)

m = mass (grams)

c = specific heat (J/g°C)

ΔT = change in temperature (°C)

2. Heat is not the same as temperature, yet they are related. Explain how they differ from each other.

Heat is a combination of kinetic energy (measured by temperature) and potential energy.

a. Perform calculations using: ($q = m \cdot c \cdot \Delta T$)

1. Gold has a specific heat of 0.129 J/(g°C). How many joules of heat energy are required to raise the temperature of 15 grams of gold from 22 °C to 85 °C?

$q = ?$

$m = 15 \text{ g}$

$c = 0.129 \text{ J/g}^\circ\text{C}$

$\Delta T = 85^\circ\text{C} - 22^\circ\text{C} = 63^\circ\text{C}$

$q = mc \Delta T$

$q = (15 \text{ g})(0.129 \text{ J/g}^\circ\text{C})(63^\circ\text{C})$

$q = 121.9 \text{ J} \rightarrow \underline{120 \text{ J}}$

Endothermic or exothermic? Endothermic

3. If the temperature of 34.4 g of ethanol increases from 25 °C to 78.8 °C, how much heat has been absorbed by the ethanol? The specific heat of ethanol is 2.44 J/(g°C)

$q = ?$

$m = 34.4 \text{ g}$

$c = 2.44 \text{ J/g}^\circ\text{C}$

$\Delta T = 78.8^\circ\text{C} - 25^\circ\text{C} = 53.8^\circ\text{C} \rightarrow 54^\circ\text{C}$

$q = (34.4 \text{ g})(2.44 \text{ J/g}^\circ\text{C})(54^\circ\text{C})$

$q = 4532.544 \text{ J} \rightarrow \underline{4500 \text{ J}}$

Endothermic or exothermic? Endothermic

5. Copper has a specific heat of 0.385 J/(g°C). A piece of copper absorbs 5000 J of energy and undergoes a temperature change from 100 °C to 200 °C. What is the mass of the piece of copper?

$q = 5000 \text{ J}$

$m = ?$

$c = 0.385 \text{ J/g}^\circ\text{C}$

$\Delta T = 200^\circ\text{C} - 100^\circ\text{C} = 100^\circ\text{C}$

$$m = \frac{q}{c \Delta T}$$

$$m = \frac{5000 \text{ J}}{(0.385 \frac{\text{J}}{\text{g}^\circ\text{C}})(100^\circ\text{C})}$$

$m = 129.87 \text{ g} \rightarrow \underline{100 \text{ g}}$

Endothermic or exothermic? Endothermic

b. Determine if it's endothermic or exothermic

2. An unknown substance with a mass of 100 grams absorbs 1000 J while undergoing a temperature increase of 15 °C. What is the specific heat of the substance?

$q = 1000 \text{ J}$

$m = 100 \text{ g}$

$c = ?$

$\Delta T = 15^\circ\text{C}$

$$c = \frac{q}{m \Delta T}$$

$$c = \frac{1000 \text{ J}}{(100 \text{ g})(15^\circ\text{C})}$$

$c = 0.6666667 \text{ J/g}^\circ\text{C} \rightarrow \underline{0.7 \text{ J/g}^\circ\text{C}}$

Endothermic or exothermic? Endothermic

4. Graphite has a specific heat of 0.709 J/(g°C). If a 25 gram piece of graphite is cooled from 35 °C to 18 °C, how much energy was lost by the graphite?

$q = ?$

$m = 25 \text{ g}$

$c = 0.709 \text{ J/g}^\circ\text{C}$

$\Delta T = 18^\circ\text{C} - 35^\circ\text{C} = -17^\circ\text{C}$

$q = (25 \text{ g})(0.709 \text{ J/g}^\circ\text{C})(-17^\circ\text{C})$

$q = -301.325 \text{ J} \rightarrow \underline{-300 \text{ J}}$

Endothermic or exothermic? Exothermic

6. 45 grams of an unknown substance undergoes a temperature increase of 38 °C after absorbing 4172.4 Joules. What is the specific heat of the substance? Look at the table on page 513 of your book, and identify the substance.

$q = 4172.4 \text{ J}$

$m = 45 \text{ g}$

$c = ?$

$\Delta T = 38^\circ\text{C}$

$$c = \frac{q}{m \Delta T}$$

$$c = \frac{4172.4 \text{ J}}{(45 \text{ g})(38^\circ\text{C})}$$

$c = 2.44 \text{ J/g}^\circ\text{C} \rightarrow \underline{2.4 \text{ J/g}^\circ\text{C} \rightarrow \text{Ethanol}}$

Endothermic or exothermic? Endothermic

Worksheet Calculations Involving Specific Heat

A. Kayode Coker



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Drying Technologies in Food Processing Xiao Dong Chen, Arun S. Mujumdar, 2009-03-16 Drying is by far the most useful large scale operation method of keeping solid foods safe for long periods of time and is of fundamental importance in most sectors of food processing Drying operations need to be precisely controlled and optimized in order to produce a good quality product that has the highest level of nutrient retention and flavor whilst maintaining microbial safety This volume provides an up to date account of all the major drying technologies employed in the food industry and their underlying scientific principles and effects Various equipment designs are classified and described The impact of drying on food properties is covered and the micro structural changes caused by the process are examined highlighting their usefulness in process analysis and food design Key methods for assessing food properties of dried products are described and pre concentration and drying control strategies are reviewed Thermal hazards and fire explosion detection and prevention for dryers are discussed in a dedicated chapter Where appropriate sample calculations are included for engineers and technologists to follow The book is directed at food scientists and technologists in industry and research food engineers and drying equipment manufacturers Proceedings American Society for Engineering Education. Conference, 1993

Sustainable Building - Design Manual, 2004-01-01 This first volume of Sustainable building design manuals focuses on policy and regulatory mechanisms and serves as a guide to policy makers and local authorities **Solar Air Systems** Robert Hastings, 2013-10-18 Active solar systems for air heating are a straightforward yet effective way of using solar energy for space heating and tempering ventilation air They offer some unique advantages over solar water systems can offer improved comfort and fuller use of solar gains than passive solar systems and are a natural fit with mechanically ventilated buildings They can be economical with short pay back periods and can act not only as space heating or ventilation air heating

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