



OWNER'S MANUAL

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MINIMUM RECOMMENDED OPERATOR AGE: 16

03

TRX400EX
Sportrax400EX

Trx400ex Sportrax400ex Year 2003 Owners Manual

S Ben Porath



Trx400ex Sportrax400ex Year 2003 Owners Manual:

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Honda Trx300ex & Trx400ex Atv's 1993 Thru 2004 Mike Stubblefield, John Harold Haynes, 2005-09-01 Each Haynes Manual is based on a complete teardown and rebuild of the specific vehicle Features hundreds of hands on photographs taken of specific repair procedures in progress Includes a full chapter on scheduled owner maintenance and devotes a full chapter to emissions systems Wiring diagrams are featured throughout **Honda TRX400 Foreman 1995-2003** Penton Staff, 2000-05-24 TRX400FW 1995 2003 **Honda TRX300EX & TRX400EX and TRX405R/ER ATVs** Ken Freund, Mike Stubblefield, 2007-02-01 Haynes offers the best coverage for cars trucks vans SUVs and motorcycles on the market today Each manual contains easy to follow step by step instructions linked to hundreds of photographs and illustrations Included in every manual troubleshooting section to help identify specific problems tips that give valuable short cuts to make the job easier and eliminate the need for special tools notes cautions and warnings for the home mechanic color spark plug diagnosis and an easy to use index

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for the nitration 4.01 ml of nitric acid 4.00 ml of sulfuric acid and 4.56 ml bromobenzene were used to complete the reaction with these measurements bromobenzene was found to be the limiting reagent since only 4.56 ml of bromobenzene was used the amount of product predicted to be recovered was 8.7533 g

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the aim of the experiment is to determine and study electrophilic aromatic substitution reaction e.g. particularly nitration of methyl benzoate to give methyl m-nitrobenzoate thus determine the weight, melting point and percentage yield of the pure methyl m-nitrobenzoate abstract

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in this experiment we will carry out the nitration of bromobenzene. Bromine substituted on an aromatic ring is an ortho/para directing group. It will direct an incoming electrophile to the ortho and para positions. It is also deactivating, meaning that its reactions will be slower than the identical reactions of benzene.

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section 20711 lab experiment nitration of bromobenzene introduction the chemical reaction of nitration consists of a nitro group being added to or substituted in a molecule. Nitration can basically be carried out by a mixture of concentrated nitric acid and sulfuric acid. This mixture is useful to obtain the active nitronium ion.

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in the nitration of bromobenzene a mixture of sulfuric acid and nitric acid react with one another. Nitric acid, the weaker of the two acids, will serve as base to form nitronium ion. Bromobenzene will attack the nitronium ion by nucleophilic attack.

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then actual is your actual findings yield actual theoretical $\times 100$. What are the possible side reactions of nitration of

bromobenzene dinitration can occur at ortho para or at both ortho positions occurs in overheating meta product happens very rarely because bromine is an ortho para director

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the bromine substituent in bromobenzene is an ortho para director the para position is favored due to steric effects in the reaction that takes place in this lab a nitro group is substituted onto bromobenzene at the para position to form 1-bromo-4-nitrobenzene

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over 5 minutes slowly pipet this bromobenzene into the icy cold nitration mixture HNO_3 3 h 2 so 4 in the 125 ml flask and stir continue to stir the reaction mixture in the ice bath for 30 minutes do not heat in a new 250 ml erlenmeyer

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on the right the arenium cation has a nitro set with incredibly small 0.6 angstrom bond distances this will result in an optimization failure as the atoms are even close for the programme to accommodate their electrons demonstration of creating the p arenium cation by nitration of bromobenzene

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the purpose of this lab was to observe where substituents are added on a benzene ring by way of electrophilic aromatic substitution the electrophilic source in the nitration of bromobenzene comes from a nitronium ion NO_2^+ which acts as the

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