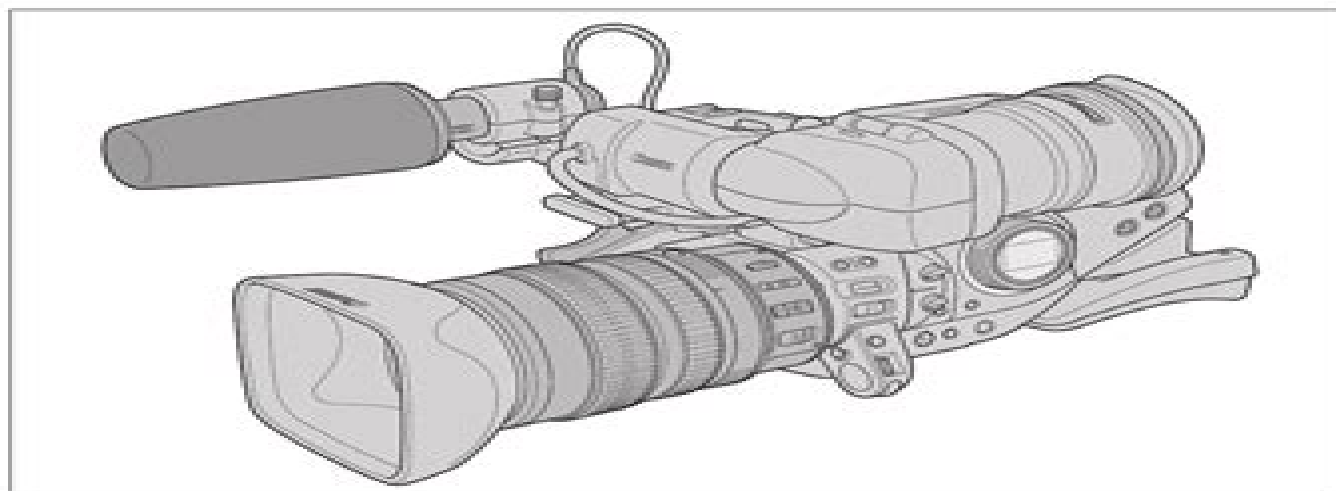


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 $\text{N}_2 + 3\text{H}_2 \rightarrow 2\text{NH}_3$
 $\text{Mg}(\text{OH})_2 + 2\text{HNO}_3 \rightarrow \text{Mg}(\text{NO}_3)_2 + 2\text{H}_2\text{O}$
 $\text{Mg} + 2\text{HCl} \rightarrow \text{MgCl}_2 + \text{H}_2$
 $2\text{Al} + 3\text{Cl}_2 \rightarrow 2\text{AlCl}_3$
 $2\text{Al} + 3\text{S} \rightarrow \text{Al}_2\text{S}_3$
 $2\text{Al} + 3\text{O}_2 \rightarrow \text{Al}_2\text{O}_3$
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 $\text{Mg}(\text{OH})_2 + \text{HCl} \rightarrow \text{MgCl}_2 + \text{H}_2\text{O}$
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 $\text{Mg} + 2\text{HCl} \rightarrow \text{MgCl}_2 + \text{H}_2$

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 $2\text{C}_2\text{H}_2 + 5\text{O}_2 \rightarrow 4\text{CO}_2 + 2\text{H}_2\text{O}$
 if 1.0 mol of C_2H_2 is burned how many moles of CO_2 are formed
 $2.5 \text{ mol C}_2\text{H}_2 \times \frac{4 \text{ mol CO}_2}{2 \text{ mol C}_2\text{H}_2} = 5.0 \text{ mol CO}_2$
 if 1.0 mol of C_2H_2 is burned how many moles of oxygen gas are consumed
 $1.0 \text{ mol C}_2\text{H}_2 \times \frac{5 \text{ mol O}_2}{2 \text{ mol C}_2\text{H}_2} = 2.5 \text{ mol O}_2$

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