

CHEMISTRY Terminology

The ABCs of chemistry

A

A: Symbol for atomic mass number
absolute zero: Lowest temperature; 0 Kelvin
Ae: Symbol for the element actinium, $Z = 89$
accuracy: The closeness of measurement to true value
acetate anion: $C_2H_3O_2^-$, charge -1
acetic acid: $C_2H_4O_2$
acetone: Dimethyl ketone; organic solvent
acid-base indicator: A chemical that changes color if the pH changes
acid-base titration: A method for determining acid or base concentrations
acid definition: See Arrhenius acid, Brønsted-Lowry acid, Lewis acid
acid ionization constant (K_a): Equilibrium constant for acid dissociation
acidic solution: pH below 7
actinide: Element with $Z = 90$ or above; radioactive
activation energy (E_a): A process's energy barrier



adhesion: Attraction of unlike molecules or materials
Ag: Symbol for the element silver, $Z = 47$
Al: Symbol for the element aluminum, $Z = 13$
alcohol: Organic compound with -OH group; ROH
aldehyde: Carbonyl group bonded to hydrogen and an organic group
alkali metal family: Lithium, sodium, potassium, rubidium, cesium, and francium; column #1
alkaline earth metal family: Beryllium, magnesium, calcium, strontium, barium, and radium; column #2
alkane: Hydrocarbon; all C-C single bonds
alkene: Hydrocarbon; 1 or more C=C double bonds
alkyne: Hydrocarbon; 1 or more C≡C triple bonds
allotropes: Two or more forms of an element
alloy: Solution of 2 or more metals
alpha (α): Greek letter that denotes radioactive particle and various scientific variables
alpha particle: Helium nucleus; charge +2
Au: Symbol for the element auriferous, $Z = 79$
amalgam: Alloy of mercury with other metal
amine: Organic base; RRRN; modified ammonia
amino acid: Compound with organic acid and organic base property; forms proteins and peptides
ammonia: NH_3 , base
ammonium cation: NH_4^+ , charge +1
ammonium hydroxide: NH_4OH
amphoteric oxide: Exhibits both acid and basic properties
amu: Atomic mass unit; $1/12$ mass of C-12
anion: Ion with a negative charge
anode: Electrode that supports oxidation
antibonding MO: MO is less stable than separate AOs
AO: Abbreviation for atomic orbital

aq: Aqueous; "dissolved in water"
Ar: Symbol for the element argon, $Z = 18$
aromatic: Organic compound with a benzene ring
Arrhenius acid: Produces hydronium ion in water solution
Arrhenius base: Yields hydroxide ion in water solution
Arrhenius equation for rate constant (k):

$$k = Ae^{-E_a/RT}$$
 E_a = activation energy
arsenide anion: As^{3-} , charge -3
As: Symbol for the element arsenic, $Z = 33$
At: Symbol for the element astatine, $Z = 85$
atm: Symbol for pressure in "atmospheres"
atom: Fundamental unit of all matter
atomic mass number (A): Total number of protons and neutrons
atomic number (Z): Number of protons in the nucleus
atomic orbital: Wave motion of electrons in atoms
atomic radius: Empirical measure of atom size
atomic weight: Weighted average of natural isotopes of an element
Au: Symbol for the element gold, $Z = 79$
Aufbau principle: Guides the filling of electronic subshells of the elements
Avogadro's law: Volume is proportional to gas moles at fixed pressure and temperature
Avogadro's number: $N_A = 6.022 \times 10^{23} \text{ mol}^{-1}$

Bk: Symbol for the element berkelium, $Z = 97$
bleach: Chemical that is a strong oxidizing agent
boiling: Liquid $\rightarrow$ gas at the boiling point
boiling point elevation: A solution has a higher boiling point than a pure solvent
boiling point (T_b): Liquid-gas equilibrium, $P = 1 \text{ atm}$
boiling point of water: 100°C
bond energy: Energy held by a chemical bond
bond length: Distance between 2 atoms in a bond
bond order for Lewis model: # of bonds divided by # of bonded atoms
bond order for MO treatment: # of filled bonding MOs minus # of filled antibonding MOs
bonding electrons: Form chemical bonds in a compound
bonding MO: MO is more stable than AOs
Boyle's law: PV constant for gas, fixed T and n
Br: Symbol for the element bromine, $Z = 35$
bromide anion: Br^- , charge -1
Brønsted-Lowry acid: Proton donor
Brønsted-Lowry base: Proton acceptor
buffer: Composed of weak acid and weak base; serves to keep pH constant



B

B: Symbol for the element boron, $Z = 5$
Ba: Symbol for the element barium, $Z = 56$
balanced equation: Each side of the equation has the same number of atoms of each element and charge
balanced redox: Electron loss = electron gain
barium cation: Ba^{2+} , charge +2
base definition: See Arrhenius base, Brønsted-Lowry base, Lewis base
base ionization constant (K_b): Equilibrium constant for base dissociation
basic solution: pH over 7
battery: Source of electrical power; galvanic cell
Be: Symbol for the element beryllium, $Z = 4$
benzene: C_6H_6 ; organic solvent; aromatic ring of 6 carbon atoms
beryllium cation: Be^{2+} , charge +2
beta (β): Greek letter that denotes radioactive particle and a number of scientific variables
beta particle: Energetic electron from the nucleus; charge -1
Bh: Symbol for the element bohrium, $Z = 107$
Bi: Symbol for the element bismuth, $Z = 83$
bicarbonate anion: HCO_3^- , charge -1
bimolecular oxygen: O_2
binary compound: Formed from 2 elements
bisulfate anion: HSO_4^- , charge -1



b: Symbol for centi- , the SI prefix for $1/100$ or 10^{-2}
c: Symbol for the speed of light
°C: Celsius temperature scale
C: Symbol for the element carbon, $Z = 6$
 $C_2H_4O_2$: Acetic acid
Ca: Symbol for the element calcium, $Z = 20$
calcium cation: Ca^{2+} , charge +2
calcium fluoride: CaF_2
calcium hydroxide: $Ca(OH)_2$
calorie: English unit of heat energy
calorimetry: Study of the release or absorption of heat
carbide anion: C^{4-} , charge -4
carbohydrate: Organic compound; carbon bonded to several -OH groups (sugar and starch)
carbon dioxide: CO_2
carbon monoxide: CO
carbonate anion: CO_3^{2-} , charge -2
carbonic acid: H_2CO_3
carbonyl group: CO , ligand complex
carboxyl group: $-CO$ bonded to 2 groups
carboxylic acid: Organic acid; $R-COOH$
catalyst: Accelerates a reaction but is not consumed in the reaction; lowers the activation energy
cathode: Electrode that supports reduction
cation: Ion with a positive charge
Cd: Symbol for the element cadmium, $Z = 48$



Thermal Chemistry Study Guide

James E. Brady, John R. Holum



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Klein, Laurie S. Starkey, 2025-03-18 Success in organic chemistry requires mastery in two core aspects fundamental concepts and the skills needed to apply those concepts and solve problems With *Organic Chemistry Student Study Guide and Solutions Manual* 5th Edition students can learn to become proficient at approaching new situations methodically based on a repertoire of skills These skills are vital for successful problem solving in organic chemistry **Chemistry, Study Guide** James E.

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