# PRACTICE TEST/THERMOCHEMISTRY/AP CHEMISTRY

3. a) Calculate the amount of heat transferred when 10.00 g of N2O(g) is formed by the following reaction:

 $2N2(q) + O2(q) \rightarrow 2N2O(q)$   $\Delta Hrxn = +163.2 kJ$ 

- b) Draw an energy diagram for this process.
- Predict the value for ΔH°f for the following scenarios:

a) Br2(g) b) Br2(1)

c) 12(9)

d) (12(5))

Calculate the AHrxn for the following reaction:

 $C_2H_4(q) + 3O_2(q) \rightarrow 2CO_2(q) + 2H_2O(l)$ 

 $\Delta H^{\circ}f = C_2H_4(q) = 226.6 \text{ kJ/mol}$  $\Delta H^{\circ}f = -393.5 \text{ kJ/mol}$  $\Delta H^{\circ}f H_2O(1) = -285.8 \text{ kJ/mol}$ 

 A 5.00 g sample of liquid water at 25.0°C is heated by the addition of 84.0 J of energy. The final temperature of the water is \_\_\_\_\_\_\_°C. The specific heat capacity of the liquid is 4.18 J/g°C.

a) -21

b) 4.02 c) 25.2

d) 95.2 e) 29.0

- 5. THIS PROBLEM WAS TAKEN DIRECTLY FROM THE 1995 AP CHEM. EXAM. Propage is a hydrocarbon that is commonly used as a fuel for cooking. Propane's formula is C3H8.
  - a) Write a balanced equation for the complete combustion of propage gas.
  - b) Calculate the volume of air at 30°C and 1.00 atm that is needed to burn completely 10.0 g of propane. Assume that air is 21.0% O2 by volume.
  - c) The heat of combustion (\(\Delta\mathbb{H}^\tilde{\combustion}\) is -2,220.1 kJ/mol. Calculate the heat of formation,  $\Delta H^{\circ}_{t}$ , of propane given that  $\Delta H_{t}$  of  $H_{2}O(1)$  is -285.3 kJ/mol and  $\Delta H^{\circ}_{t}$  of CO2(q) is -393.5 kJ/mol.
  - d) Assuming that all of the heat evolved burning 10.0 g propane is transferred to 8.00 kg of water (specific heat = 4.184 J/q°C), calculate the increase in temperature of the wwater.

# **Thermochemistry Practice Test**

Neil D. Jespersen, Pamela Kerrigan

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#### **Table of Contents Thermochemistry Practice Test**

- 1. Understanding the eBook Thermochemistry Practice Test
  - The Rise of Digital Reading Thermochemistry Practice Test
  - Advantages of eBooks Over Traditional Books
- 2. Identifying Thermochemistry Practice Test
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Thermochemistry Practice Test
  - User-Friendly Interface
- 4. Exploring eBook Recommendations from Thermochemistry Practice Test
  - Personalized Recommendations

- Thermochemistry Practice Test User Reviews and Ratings
- Thermochemistry Practice Test and Bestseller Lists
- 5. Accessing Thermochemistry Practice Test Free and Paid eBooks
  - Thermochemistry Practice Test Public Domain eBooks
  - Thermochemistry Practice Test eBook Subscription Services
  - Thermochemistry Practice Test Budget-Friendly Options
- 6. Navigating Thermochemistry Practice Test eBook Formats
  - ∘ ePub, PDF, MOBI, and More
  - Thermochemistry Practice Test Compatibility with Devices
  - Thermochemistry Practice Test Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Thermochemistry Practice Test
  - Highlighting and Note-Taking Thermochemistry Practice Test
  - Interactive Elements Thermochemistry Practice Test
- 8. Staying Engaged with Thermochemistry Practice Test
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Thermochemistry Practice Test
- 9. Balancing eBooks and Physical Books Thermochemistry Practice Test
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Thermochemistry Practice Test
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Thermochemistry Practice Test
  - $\circ\,$  Setting Reading Goals Thermochemistry Practice Test
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Thermochemistry Practice Test
  - Fact-Checking eBook Content of Thermochemistry Practice Test

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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