

Thermodynamics - Example Problems

Problems and Solutions

1 Examining a Power Plant

Consider a power plant. At point 1 the working gas has a temperature of $T = 25^\circ\text{C}$. The pressure is 1bar and the mass flow is $\dot{m} = 2\text{kg/s}$. After point 1, the air enters a compressor, at which the pressure is increased to 6 bar. (This is point 2.) After this, the flow is heated to $T_3 = 700^\circ\text{C}$. (This is stage 3.) Afterwards, the flow goes through a turbine, at which the pressure is decreased to 1bar again. (This is stage 4.) The flow is then cooled to $T_1 = 25^\circ\text{C}$. We're now back in stage 1.

We assume that:

1. air behaves like a thermally and calorically perfect gas with $C_p = 1004\text{J/kg K}$, $R = 287\text{J/kg K}$ and $\gamma = 1.4$.
2. the isentropic efficiency of the compressor is 0.88.
3. the polytropic efficiency of the turbine is 0.85.

Answer the following questions.

1. Depict the evolution of the cycle in the $T - s$ plane.
2. Determine the net power.
3. Determine the heat received from the hot source.
4. Determine the thermal efficiency of the power plant.
5. Determine the Carnot efficiency associated with the present power plant and compare it with the previous result.
6. Verify the first principle.

1 Solution

1. Let's draw the $T - s$ graph. We start by drawing two lines of constant pressure. (This can be seen in the graph below.)

We now start at point 1. At this point the pressure and temperature are low. When we go to point 2, we don't go straight up, since the process is not isentropic. Instead, we go up and slightly right as well, until we reach the line $p = 6\text{bar}$. We then have found point 2.

The heating takes place under constant pressure. So from point 2 to point 3 we stay on the isobaric line $p = 6\text{bar}$. This goes on until we reach $T = 973\text{K}$. This is the position of point 3.

The turbine isn't isentropic either. So from point 3 we don't go straight down. Instead, we go down and a bit to the right as well, until we reach the line $p = 1\text{bar}$. This is point 4.

From point 4 we return to point 1. Since the cooling occurs at constant pressure, we do this without leaving the line $p = 1\text{bar}$.

Thermodynamic Example Problems

Edenilson Brandl



Thermodynamic Example Problems:

Thermodynamics Problem Solving in Physical Chemistry Kathleen E. Murphy, 2020-03-23 Thermodynamics Problem Solving in Physical Chemistry Study Guide and Map is an innovative and unique workbook that guides physical chemistry students through the decision making process to assess a problem situation create appropriate solutions and gain confidence through practice solving physical chemistry problems The workbook includes six major sections with 20 30 solved problems in each section that span from easy single objective questions to difficult multistep analysis problems Each section of the workbook contains key points that highlight major features of the topic to remind students of what they need to apply to solve problems in the topic area Key Features Provides instructor access to a visual map depicting how all equations used in thermodynamics are connected and how they are derived from the three major energy laws Acts as a guide in deriving the correct solution to a problem Illustrates the questions students should ask themselves about the critical features of the concepts to solve problems in physical chemistry Can be used as a stand alone product for review of Thermodynamics questions for major tests *Thermodynamics 1 With No Calculations* Edenilson Brandl, Have you ever wanted to understand thermodynamics without getting lost in complex calculations This book provides a clear intuitive approach to one of the most important scientific fields making it accessible to students professionals and curious minds alike By focusing on concepts rather than equations it offers a fresh and engaging way to learn how energy works in the world around us Whether you re a beginner looking for an introduction or an expert seeking a new perspective this book will give you the tools to grasp the principles of thermodynamics with ease With real world examples simple explanations and insightful discussions you will walk away with a deeper appreciation of how thermodynamics shapes our daily lives and the technologies we rely on Don t let complex mathematics stand in the way of understanding thermodynamics Dive into this book and discover how energy heat and entropy govern everything from the engines that power our world to the fundamental processes of nature Get your copy today and unlock the secrets of thermodynamics in the simplest way possible Modern Engineering Thermodynamics Robert Balmer, 2011-01-25 Modern Engineering Thermodynamics is designed for use in a standard two semester engineering thermodynamics course sequence The first half of the text contains material suitable for a basic Thermodynamics course taken by engineers from all majors The second half of the text is suitable for an Applied Thermodynamics course in mechanical engineering programs The text has numerous features that are unique among engineering textbooks including historical vignettes critical thinking boxes and case studies All are designed to bring real engineering applications into a subject that can be somewhat abstract and mathematical Over 200 worked examples and more than 1 300 end of chapter problems provide opportunities to practice solving problems related to concepts in the text Provides the reader with clear presentations of the fundamental principles of basic and applied engineering thermodynamics Helps students develop engineering problem solving skills through the use of structured problem solving techniques Introduces the Second Law of

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Solved Problems in Mechanical Engineering Thermodynamics P. E. Liley, 1989 *Fundamentals of Engineering Thermodynamics* Michael J. Moran, Howard N. Shapiro, Daisie D. Boettner, Margaret B. Bailey, 2018-05-14 Fundamentals of Engineering Thermodynamics sets the standard for teaching students how to be effective problem solvers Integrated throughout this product are real world applications that emphasize the relevance of thermodynamic principles to some of the most critical problems and issues of today These include a wealth of coverage on topics related to energy and the environment biomedical bioengineering and emerging technologies The Enhanced E Text is also available bundled with an abridged print companion and can be ordered by contacting customer service here ISBN 9781119456285 Price 130 95 Canadian Price 114 95

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Edition helps readers master the fundamentals of applied thermodynamics as practiced today with extensive development of molecular perspectives that enables adaptation to fields including biological systems environmental applications and nanotechnology This text is distinctive in making molecular perspectives accessible at the introductory level and connecting properties with practical implications Features of the second edition include Hierarchical instruction with increasing levels of detail Content requiring deeper levels of theory is clearly delineated in separate sections and chapters Early introduction to the overall perspective of composite systems like distillation columns reactive processes and biological systems Learning objectives problem solving strategies for energy balances and phase equilibria chapter summaries and important equations for every chapter Extensive practical examples especially coverage of non ideal mixtures which include water contamination via hydrocarbons polymer blending recycling oxygenated fuels hydrogen bonding osmotic pressure electrolyte solutions zwitterions and biological molecules and other contemporary issues Supporting software in formats for both MATLAB and spreadsheets Online supplemental sections and resources including instructor slides ConcepTests coursecast videos and other useful resources

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