

CHEM 101 Worksheet, Chapter 6: Thermochemistry

Multiple choice questions 1-10 (page 1): circle the correct answer(A-E)

Show your work questions 11-13 (pages 2-3): give the complete answer in the space provided

R = 8.314 J/(mol.K) = 0.08206 L.atm/(mol.K) 1 cal = 4.184 J 0 °C = 273 K

1. Calculate the heat effect q (in kJ) when 112 g of iron (Fe(s)) is heated from 20 °C to 300 °C. The specific heat of Fe(s) = 0.48 J/g.K

- A. -34.9 kJ B. +34.9 kJ C. -15.0 kJ D. -270 J E. +15.0 kJ

2. The dissolution reaction for ammonium chloride: $\text{NH}_4\text{Cl(s)} \rightarrow \text{NH}_4^+(\text{aq}) + \text{Cl}^-(\text{aq})$ is endothermic. When ammonium chloride is dissolved in water in an insulated (adiabatic) system, the temperature will:

- A. increase B. decrease C. stay the same

3. Consider the reaction: $\text{CaCO}_3(\text{s}) \rightarrow \text{CaO(s)} + \text{CO}_2(\text{g})$. $\Delta H_{\text{rxn}} = +178 \text{ kJ}$ What is the enthalpy change when 20.0 kg CaO is formed?

- A. $+6.36 \times 10^4 \text{ kJ}$ B. $-6.36 \times 10^4 \text{ kJ}$ C. $+6.36 \times 10^4 \text{ J}$ D. 63.6 kJ E. +178 kJ

4. Consider the reaction: $2\text{NO} + \text{O}_2 \rightarrow 2\text{NO}_2$. $\Delta H_{\text{rxn}} = -114.6 \text{ kJ}$ As always, the ΔH is "for the reaction as written"! Calculate the enthalpy change ΔH , when 12.6 g NO_2 is produced.

- A. -31.4 kJ B. +31.4 kJ C. -15.7 kJ D. +15.7 kJ E. none of these

5. How much heat is associated with the reaction that produces 256.1 g of ammonia, NH_3 ?



- A. 92.45 kJ of heat is released B. 696.1 kJ of heat is released C. 184.9 kJ of heat is absorbed
D. 184.9 kJ of heat is released E. 92 kJ of heat is released

6. Which two of the following are not formation reactions?

1. $\frac{1}{2}\text{N}_2(\text{g}) + \frac{1}{2}\text{H}_2(\text{g}) \rightarrow \text{NH}_3(\text{g})$ 2. $\text{C(s)} + \text{O}_2(\text{g}) \rightarrow \text{CO}_2(\text{g})$ 3. $2\text{SO}_2(\text{g}) + \text{O}_2(\text{g}) \rightarrow 2\text{SO}_3(\text{g})$
4. $2\text{Fe(s)} + \frac{1}{2}\text{O}_2(\text{g}) \rightarrow \text{Fe}_2\text{O}_3(\text{s})$ 5. $\frac{1}{2}\text{H}_2(\text{g}) + \frac{1}{2}\text{I}_2(\text{g}) \rightarrow \text{HI(g)}$

- A. 1 and 2 B. 3 and 4 C. 2 and 3 D. 4 and 5 E. 3 and 5

7. Consider the thermochemical reaction equation: $2\text{CO(g)} + \text{O}_2(\text{g}) \rightarrow 2\text{CO}_2(\text{g})$ $\Delta H = -198 \text{ kJ}$

What is the enthalpy change for the reaction: $\text{CO}_2(\text{g}) \rightarrow \text{CO(g)} + \frac{1}{2}\text{O}_2(\text{g})$

- A. +198 kJ B. -99 kJ C. +99 kJ D. +396 kJ E. -198 kJ

8. Calculate the enthalpy change for the reaction: $2\text{KClO}_3 \rightarrow 2 \text{KCl} + \text{O}_2(\text{g})$

Standard enthalpy of formation values: $\Delta H_f^\circ(\text{KClO}_3) = -389 \text{ kJ/mol}$ $\Delta H_f^\circ(\text{KCl}) = -436 \text{ kJ/mol}$

- A. -94 kJ B. -47 kJ C. +94 kJ D. +47 kJ
E. Cannot be calculated because $\Delta H_f^\circ(\text{O}_2(\text{g}))$ is not given

9. Given the following two thermochemical equations:



What is the enthalpy change for the reaction: $2\text{H}_2\text{O}_2(\text{l}) \rightarrow 2\text{H}_2\text{O(l)} + \text{O}_2(\text{g})$?

- A. -98 kJ B. +98 kJ C. +196 kJ D. -196 kJ E. -668 kJ

10. Which of the following statements is correct for a combustion reaction in an adiabatic "bomb" calorimeter?

- A. the pressure is constant B. the temperature is constant C. the volume is constant
D. the calorimeter constant's C units are J/g.K E. $\Delta T(\text{K}) = \Delta T(^{\circ}\text{C}) + 273.15$

Show your work for the following questions.

Thermochemistry Multiple Choice

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A Textbook of Physical Chemistry, 6th Edition Sharma K.K. & Sharma L.K., 2016 A Textbook of Physical Chemistry **IIT Chemistry-II**,

Psychometrics R. Michael Furr, 2021-08-13 In this fully revised Fourth Edition of Psychometrics An Introduction author R Michael Furr centers his presentation around a conceptual understanding of psychometric core issues such as scales reliability and validity Focusing on purpose rather than procedure and the why rather than the how to this accessible book uses a wide variety of examples from behavioral science research so readers can see the importance of psychometric fundamentals in research By emphasizing concepts logic and practical applications over mathematical proofs this book gives students an appreciation of how measurement problems can be addressed and why it is important to address them The book offers readers the most contemporary views of topics in psychometrics available in the nontechnical psychometric literature

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Table of Contents Thermochemistry Multiple Choice

1. Understanding the eBook Thermochemistry Multiple Choice
 - The Rise of Digital Reading Thermochemistry Multiple Choice
 - Advantages of eBooks Over Traditional Books
2. Identifying Thermochemistry Multiple Choice
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Thermochemistry Multiple Choice
 - User-Friendly Interface
4. Exploring eBook Recommendations from Thermochemistry Multiple Choice
 - Personalized Recommendations
 - Thermochemistry Multiple Choice User Reviews and Ratings
 - Thermochemistry Multiple Choice and Bestseller Lists

5. Accessing Thermochemistry Multiple Choice Free and Paid eBooks
 - Thermochemistry Multiple Choice Public Domain eBooks
 - Thermochemistry Multiple Choice eBook Subscription Services
 - Thermochemistry Multiple Choice Budget-Friendly Options
6. Navigating Thermochemistry Multiple Choice eBook Formats
 - ePub, PDF, MOBI, and More
 - Thermochemistry Multiple Choice Compatibility with Devices
 - Thermochemistry Multiple Choice Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Thermochemistry Multiple Choice
 - Highlighting and Note-Taking Thermochemistry Multiple Choice
 - Interactive Elements Thermochemistry Multiple Choice
8. Staying Engaged with Thermochemistry Multiple Choice
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Thermochemistry Multiple Choice
9. Balancing eBooks and Physical Books Thermochemistry Multiple Choice
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Thermochemistry Multiple Choice
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Thermochemistry Multiple Choice
 - Setting Reading Goals Thermochemistry Multiple Choice
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Thermochemistry Multiple Choice
 - Fact-Checking eBook Content of Thermochemistry Multiple Choice
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

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- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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