

Chemistry 30 Thermochemistry Quiz #2 Practice

Use the following information to answer the next question:

The combustion of ethanol ($\text{C}_2\text{H}_5\text{OH}$) at 298 K, standard state ($\text{C}_2\text{H}_5\text{OH}(l)$), can be represented by the following thermochemical equation:



Equation:



$$\Delta H = -1366.8 \text{ kJ}$$



$$\Delta H = -455.6 \text{ kJ}$$

4. The enthalpy change for the combustion of ethanol as shown in the thermochemical equation is:

- a. -1366.8 kJ
- b. -455.6 kJ
- c. -1366.8 kJ
- d. -455.6 kJ

Use the following information to answer the next question:

When carbon is completely combusted (combustion), it represents the following equation:



5. The maximum heat enthalpy information with the reaction is

- a. -393.5 kJ/mol
- b. -393.5 kJ/mol
- c. -393.5 kJ/mol
- d. -393.5 kJ/mol

Thermochemistry Question Answer Guide

Disha Experts

A decorative graphic element consisting of a light blue horizontal bar with a rounded right end, and a red circular shape partially visible behind it.

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