

1a) When ever a metal is dipped in its own salt solution it has capacity to either loose or gain electrons, due to which it gains some potential which is known as single electrode potential & is represented as 'E'

Nernst equation gives the relation ship between the electrode potential concentration of the species and was derive by a well known scientist Nernst. it was derived by taking well known thermodynamic considerations

Consider a reversible equation



The decrease in free energy is given by Vant hoffs reaction isotherm

$$\Delta G = \Delta G^\circ + RT \ln k \quad 1$$

ΔG = Gibbs free energy ΔG° = Standard gibbs free energy

R = Gas constant T = Temperature k = equilibrium constant

Decrease in free energy represents the maximum work done

$$\text{i.e., } -\Delta G = W_{\max}$$

$$\text{i.e., } W_{\max} = nFE$$

where 'n' is the total number of electrons lost during oxidation or gained during reduction

$$\text{therefore } \Delta G = -nFE \quad 2$$

$$\Delta G^\circ = -nFE^\circ \quad 3$$

k = $\frac{\text{concentration of products}}{\text{concentration of reactants}}$

Substituting 2 & 3 in 1 we get,

$$-nFE = -nFE^\circ + RT \ln \frac{[M]}{[M^{n+}]}$$

Divide the whole equation by $-nF$ we get,

$$E = E^\circ - \frac{RT}{nF} \ln \frac{[M]}{[M^{n+}]}$$

Converting $\ln$ to $\log$ & changing the sign we get

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Albert A Gayle



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Table of Contents Vtu Dipmaths 1question Paper

1. Understanding the eBook Vtu Dipmaths 1question Paper
 - The Rise of Digital Reading Vtu Dipmaths 1question Paper
 - Advantages of eBooks Over Traditional Books
2. Identifying Vtu Dipmaths 1question Paper
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Vtu Dipmaths 1question Paper
 - User-Friendly Interface
4. Exploring eBook Recommendations from Vtu Dipmaths 1question Paper
 - Personalized Recommendations
 - Vtu Dipmaths 1question Paper User Reviews and Ratings
 - Vtu Dipmaths 1question Paper and Bestseller Lists

5. Accessing Vtu Dipmaths 1question Paper Free and Paid eBooks
 - Vtu Dipmaths 1question Paper Public Domain eBooks
 - Vtu Dipmaths 1question Paper eBook Subscription Services
 - Vtu Dipmaths 1question Paper Budget-Friendly Options
6. Navigating Vtu Dipmaths 1question Paper eBook Formats
 - ePub, PDF, MOBI, and More
 - Vtu Dipmaths 1question Paper Compatibility with Devices
 - Vtu Dipmaths 1question Paper Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Vtu Dipmaths 1question Paper
 - Highlighting and Note-Taking Vtu Dipmaths 1question Paper
 - Interactive Elements Vtu Dipmaths 1question Paper
8. Staying Engaged with Vtu Dipmaths 1question Paper
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Vtu Dipmaths 1question Paper
9. Balancing eBooks and Physical Books Vtu Dipmaths 1question Paper
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Vtu Dipmaths 1question Paper
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Vtu Dipmaths 1question Paper
 - Setting Reading Goals Vtu Dipmaths 1question Paper
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Vtu Dipmaths 1question Paper
 - Fact-Checking eBook Content of Vtu Dipmaths 1question Paper
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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