

- Men: Intet "formet condictio indebit" -begi
 hvis betalingen skyldes en kriminell handling
- Unntak: Betaling av foreidet gjeld kan ikke kreves tilbake
- Utgangspunkt: Tilbakebetalingsplikt?
- -Tja. Bergsåker → Nei (unt: tillegysbetaling). Lilleholt → Ja.
- Vurderingstema: Konkret helhetsvurdering av hva som er mest "rimelig", 85/290 (Birgo)
- Motstående hensyn: Oppgjørshensynet VS korreksjonshensynet.
- · Momenter:
- Skyld på noen en av sidene? Ond/god tro?
- Objektiv årsak: Forhold på betalers/mottakers side?
- Graden av betalingspress? (to mer press, jo mer rimelig med restitusjon)
- Tiden som har gått: Lang eller kort tid? (to lengre tid → äke restitusjon)
- Betalingens art: Kontanthensynet; Behov for investible (ex for mye løne)
- Mottakers konkrete innrettelse (to mer innrettelse + ikke restitusjon)
- Partenes stilling: Profesjonelle parter / forbruker
- Hvis tatt forbehold om tilbakebetaling ved feil Ofte avgjørende.

 Avtalefrihet: Partene kan alltids bli enige om endelig oppgigr eller tilbakebetaling

-Unt: Skal en god del til hvis det først er

Hvis forbrukt: Berikelsesavleggelse?

-Antakelig ja (Indebitum betyr tross alt

rimelig med tilbakebetaling

berikelse...)



 Itke konsekvent lovgivning. Reglene er i hovedsak ulovfestede

(3) AOP?

- · Jaul
- Dommer: Rt 1985 s. 290 (Birgo), Rt 1997 s.
 1029 (M/S Marlin), Rt 2001 s. 1580 (Slåtto Husbygg), Rt 2010 s. 816 (Konsesjonsavgift),

HR-2016-2491-A

J. C. Sonnekus

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web solution w d 1 5936 0 500mol 8 3145jk1mol1 300k d 1 99 103j q d w d 1 99 103j 3 5 this problem is designed to test the assertion on page 60 that for typical thermodynamic pro cesses in which the elevation of the center of mass changes it is usually a good approximation to set w equal to wlab

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web solution for a refrigerator cop low t high t low solve for the hot side temperature 45 c low 273 high t low cop 10 45 c 273 250 8k answer is a 2 helium

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web question 1 what is true of isothermal process a Δt 0 b Δu 0 c Δq Δw d pv constants solution in an isothermal process temperature remains constant Δt 0 since internal energy depends on the temperature Δu 0 from first law of thermodynamics Δu Δq Δw since Δu 0 Δq Δw also pv nrt as t is constant pv