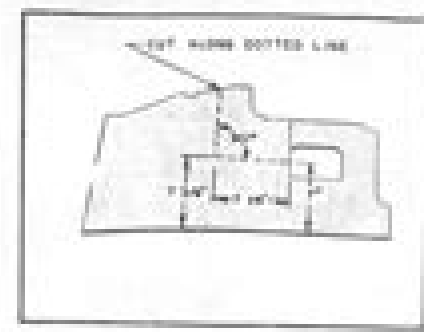
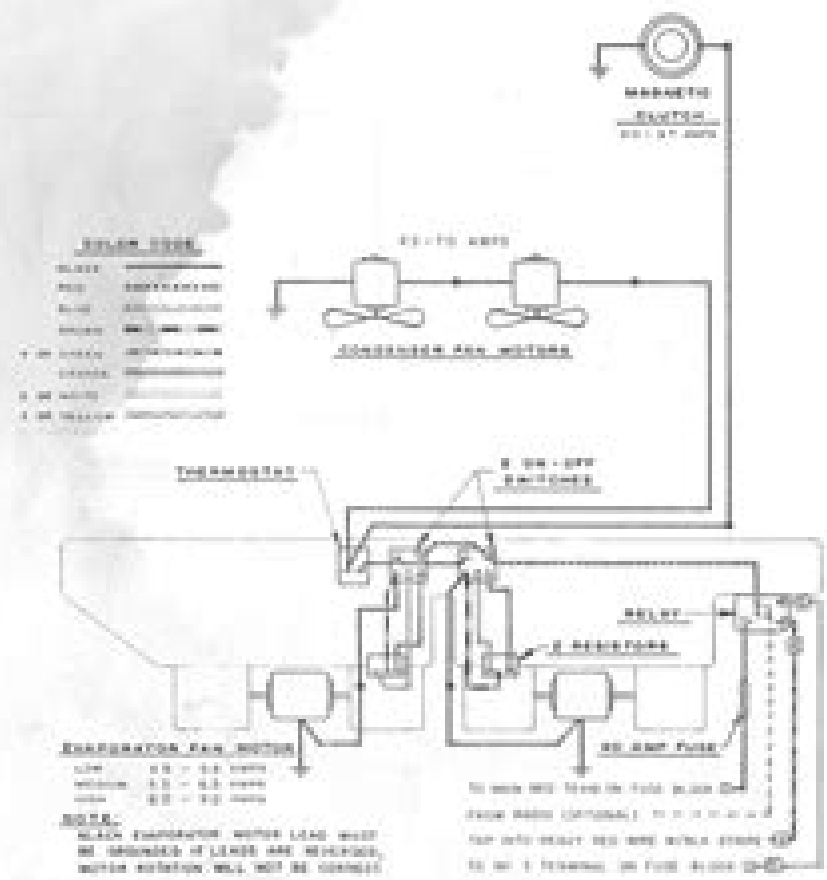
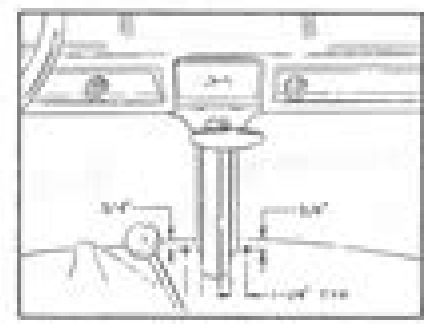


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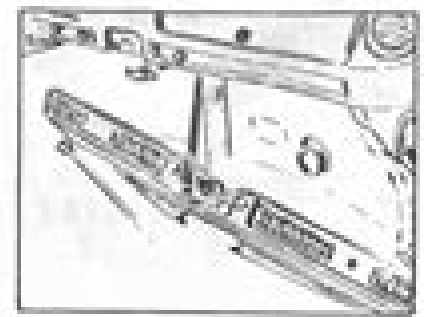
1.1 Remove and retain black panel and hardware on driver's side of vehicle. Modify panel as illustrated.



1.2 Remove and retain hardware securing fuse block and let hang.

1.3 Locate and drill (see 1/8" dia.) drain hole hole in floor panel as illustrated. CAUTION: DO NOT DRILL THROUGH CARPET.

1.4 Use special key of wire extensions to extend wire on fuse block. NOTE: EXTEND WIRE FOR WIRE BUNDLE AND COLOR CODE.



1.5 Remove white wire with black stripe from 3.5.

1.6 Modify wire harness and secure from floor holder at steering column and move to left and right of column under plastic cover. This will allow clearance for the evaporator case to fit properly.

1.7 Position evaporator case on vehicle floorboard as illustrated.

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J Ma



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l when the half reactions listed above are combined into a balanced chemical equation the ratio of the coefficients $\text{Cu} : \text{MnO}_4^{-} : 4\text{H}^{+}$ is

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web balance each half reaction in basic solution $4\text{SO}_3^{2-} + 2\text{MnO}_2 + \text{Cr}_3 + \text{NO}_3 \rightarrow \text{SO}_2 + \text{Mn}_2\text{O}_3$ balance each redox reaction in acid solution using the half reaction method $8\text{H}_2\text{O}_2 + \text{Cr}_2\text{O}_7^{2-} + 2\text{H}^+ \rightarrow 2\text{H}_2\text{O} + 2\text{Cr}^{3+} + 10\text{H}_2\text{O}$ $11\text{PbO}_2 + \text{IO}_3^- + 2\text{H}^+ \rightarrow \text{PbO}_2 + \text{IO}_3^- + \text{H}_2\text{SO}_4$

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web practice exercises balance the following equations of redox reactions assign oxidation numbers to all elements in the reaction separate the redox reaction into two half reactions balance the atoms in each half reaction add the two half reactions together and cancel out common terms

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web aug 29 2023 some points to remember when balancing redox reactions the equation is separated into two half equations one for oxidation and one for reduction the equation is balanced by adjusting coefficients and adding H^+ and OH^- in this order balance the atoms in the equation apart from o and h

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web in the redox equation below that occurs in our stomach which of the following pairs identifies the reducing and oxidizing agents respectively $\text{C}_6\text{H}_8\text{O}_6 + 2\text{HNO}_2 \rightarrow \text{C}_6\text{H}_6\text{O}_6 + 2\text{H}_2\text{O} + 2\text{NO}$ choose 1 answer $\text{C}_6\text{H}_8\text{O}_6$ and NO_2 a $\text{C}_6\text{H}_8\text{O}_6$ and NO_2 b NO_2 and $\text{C}_6\text{H}_8\text{O}_6$ c $\text{C}_6\text{H}_8\text{O}_6$ and HNO_2 d HNO_2 and $\text{C}_6\text{H}_8\text{O}_6$