

# Stoichiometry Study Guide

Balance the following equation. Then use the balanced equation to answer the questions below.

11. If there are  
a) How many

- b) How many

- c) How many

12. If there are  
a) How many

- b) How many

- c) How many

13. If there are  
a) How many

- b) How many

- c) How many

Name \_\_\_\_\_ Period \_\_\_\_\_ Date \_\_\_\_\_

## Stoichiometry Study Guide

All in the blanks below:

1 mole = \_\_\_\_\_ grams = \_\_\_\_\_ particles = \_\_\_\_\_ liters

Balance the following equation. Then use the balanced equation to answer the questions below.



11. If there are 12 moles of  $\text{O}_2$  used in the reaction:

a) How many grams of NaCl would be needed to react?

b) How many grams of NaCl would be needed to react?

c) How many grams of NaCl would be needed to react?

d) How many grams of NaCl would be needed to react?

12. If the above chemical reaction is done as an experiment using 10 grams of each reactant, calculate how much NaClO<sub>2</sub> will be produced.

a) Given 10 grams of NaCl, how many moles of NaClO<sub>2</sub> will be produced?

b) Given 10 grams of  $\text{O}_2$ , how many moles of NaClO<sub>2</sub> will be produced?

c) Which reactant is the limiting reactant? Which reactant is the excess reactant?

13. If the above chemical reaction is done as an experiment using 10 grams of each reactant, calculate the percent yield given that the actual yield of NaClO<sub>2</sub> is 10.8 grams.

a) What is the theoretical yield in grams of NaClO<sub>2</sub> given 10 grams of NaCl?

b) What is the percent yield given that the actual yield of NaClO<sub>2</sub> is 10.8 grams?

c) What is the theoretical yield in grams of NaClO<sub>2</sub> given 10 grams of  $\text{O}_2$ ?

d) What is the percent yield given that the actual yield of NaClO<sub>2</sub> is 10.8 grams?

# Unit 11 Stoichiometry Study Guide Answer Key

**Mun Prasad**



## **Unit 11 Stoichiometry Study Guide Answer Key:**

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