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1 st Law of Thermodynamics:	Internal Energy:
Endothermic: $q=+\ell$ Exothermic: $q=-\ell$	$\Delta U = q + W$
Work Done By/On a Gas:	$W = -P\Delta V$
Work done by the system: W=−i Work done on the system: W=+i	1 L +1 otm = 101.3 Joules
Specific Heat Capacity: J/(g&°C)&	$C_s = \frac{q}{m \Delta T}$
Molar Heat Capacity: J/(mol °C)	$C_{\omega} = \frac{q}{n\Delta T}$
Heat Capacity: 1/°C	$C_{N} = \frac{q}{\Delta T}$
Calorimetry - Temperature Change:	$q = mc \Delta T$ $C_{H20} = 4.184 \frac{J}{\alpha^{9}C} C_{tot} = 2.09 \frac{J}{\alpha^{9}C} C_{3res} = 2.03 \frac{J}{\alpha^{9}C}$
Calorimetry - Phase Change:	$q=m\Delta H q=n\Delta H$
	H2O: ΔH_{pas} =334 $J/g \Delta H_{pas}$ =6.01 kJ/mol H2O: ΔH_{sap} =2260 $J/g \Delta H_{sap}$ =40.7kJ/mol
Final Temperature of a Mixture:	$-q_A = q_B$
T▲ → Initial Temp. of Substance A	$-m_{A}C_{A}[T_{F}-T_{A}]=m_{0}C_{0}(T_{F}-T_{0})$
T _B → Initial Temp. of Substance B	Note: This equation will work if there are no phase changes.
Coffee Cup Calorimeter:	$\Delta H^{\circ}_{cos} = \frac{q_{cos}}{n} q_{cos} = -q_{eq} q_{eq}$
Enthalpy Change:	

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