

1. Prepping the crucibles and crucible lids. Observe, clean and dry two crucibles and their lids. As mentioned in the safety precautions, the crucibles will get hot when heated and feel changes color to marooned in the safety precautions, the crucibles with their lids on *you* use clay triangles. The lid should be opened hot. Using tongs, place the crucibles with their lids on *you* use clay triangles. Then heat the crucibles with their lids slightly to allow water vapor to easily escape the crucible. Then heat the crucibles with their lids off for about four minutes. The point of this is to make sure the crucibles and their lids are completely dry. Turn the burner off and wait three minutes. **THE CRUCIBLES ARE STILL HOT!** Using tongs, carry the crucibles to your bench. **THE CRUCIBLES ARE STILL HOT!** Wait 10 minutes while the crucibles and lids further cool. While the crucibles and lids are cooling proceed to step 2.
2. Obtain 1.0 grams of each hydrate and place each in a separate crucible, after the crucibles have cooled. $\text{CuSO}_4 \cdot x \cdot \text{H}_2\text{O}$ is the blue sample, while $\text{CoCl}_2 \cdot x \cdot \text{H}_2\text{O}$ is the purple sample. Put the crucible lids on top of the crucibles. Weigh each crucible, its lid and the hydrate, record these values in the data table at the top of page 3. In each sample x is an unknown integer.
3. Place each crucible with its lid on a clay support triangle and heat each for five minutes on low flame. After this slightly increase the flame and continue heating for ten minutes. After this time, turn off the burner and wait three minutes. **THE CRUCIBLES ARE STILL HOT!** Using tongs, carry the crucible burner and wait 10 minutes for the crucible to cool. Increasing the temperature slowly is to minimize the probability of spattering of the sample during dehydration. The lid is to prevent the loss of sample in case of some spattering. The cobalt chloride hydrate often does spatter if heated rapidly. Lastly, do not let the crucible get red hot since unwanted sample decomposition may occur.
4. Carry the crucible to the balance with tongs, the crucibles still will be slightly hot. Weigh each crucible lid and now anhydrous sample. Record these values.
5. Repeat steps three and four. Our goal here is to drive off any remaining water in the sample. If the difference in the masses between the two measurements is greater than 0.01 g, steps three and four must be repeated yet again. Record these values and all values in Table 1.
6. Keep a note of the colors of the samples before and after dehydration.

TABLE 1
DATA TABLE

Parameter	$\text{CuSO}_4 \cdot x \cdot \text{H}_2\text{O}$	$\text{CoCl}_2 \cdot x \cdot \text{H}_2\text{O}$
Mass of crucible and lid (g)	N/A	N/A
Mass of crucible, lid and sample (g)	N/A	N/A
Mass of sample (g)	1.000 g	1.000 g
Mass of crucible, lid and sample after first heating (g)	No Entry Here	No Entry Here
after second heating (g)	0.718 g	0.731 g
Mass of water lost (g)	0.282 g	0.269 g
Percent of water in sample by mass	28.2%	26.9%

Way To Go Indigo Lab Pre Lab Answers

Michael Hoelscher



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Table of Contents Way To Go Indigo Lab Pre Lab Answers

1. Understanding the eBook Way To Go Indigo Lab Pre Lab Answers
 - The Rise of Digital Reading Way To Go Indigo Lab Pre Lab Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Way To Go Indigo Lab Pre Lab Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Way To Go Indigo Lab Pre Lab Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Way To Go Indigo Lab Pre Lab Answers
 - Personalized Recommendations
 - Way To Go Indigo Lab Pre Lab Answers User Reviews and Ratings
 - Way To Go Indigo Lab Pre Lab Answers and Bestseller Lists
5. Accessing Way To Go Indigo Lab Pre Lab Answers Free and Paid eBooks
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 - Way To Go Indigo Lab Pre Lab Answers Enhanced eBook Features
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 - Highlighting and Note-Taking Way To Go Indigo Lab Pre Lab Answers
 - Interactive Elements Way To Go Indigo Lab Pre Lab Answers
- 8. Staying Engaged with Way To Go Indigo Lab Pre Lab Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Way To Go Indigo Lab Pre Lab Answers
- 9. Balancing eBooks and Physical Books Way To Go Indigo Lab Pre Lab Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Way To Go Indigo Lab Pre Lab Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Way To Go Indigo Lab Pre Lab Answers
 - Setting Reading Goals Way To Go Indigo Lab Pre Lab Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Way To Go Indigo Lab Pre Lab Answers
 - Fact-Checking eBook Content of Way To Go Indigo Lab Pre Lab Answers
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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