3.1 Solution

Given: initial gas volume = 5.667 × 10⁻⁸ m³

initial gas temperature = 294.3 K initial gas pressure = 10.13 bar

gas doubles in volume to $1.133 \times 10^{-6} \text{ m}^3$

change in volume = $5.667 \times 10^{-6} \text{ m}^3$

Choose, as the system, the universe or surroundings exclusive of the gas in the cylinder. Then, by the constraints given, including no friction in surroundings:

$$Q = 0$$

metal.

$$-W = P_{\omega} \left(\underline{V}_{r} - \underline{V}_{r} \right) + m_{\mu} g \Delta Z$$

well-care

P_ = atmospheric pressure = 1.013 bar

 $V_r - V_r =$ change in system volume = -5.667 × 10⁻³ m³

m_p = muss of pixton = 225 kg g = acceleration due to gravity = 9.81 m/s²

AZ = variation in vertical position of piston

= -0.3048 m if pinton rises

0 if piston moves horizontally

- + 0.3048 m if piston falls

Case 1: Pistow rises (note 1 bor = 10⁵ N/m²)

$$W = -(1.013 \times 10^{5}) (-5.667 \times 10^{-5}) - (226) (9.81) (-0.3048) = +1.250 \text{ kJ}$$

$$\Delta E_{\mathrm{summ}} = Q + W = 1.250 \; \mathrm{kJ} = -\Delta E_{\mathrm{geo}} = -N_{_{Z}} C_{_{X}} (T_{_{X_{\mathrm{c}}}} - T_{_{X}})$$

$$N_e$$
 = moles of gas = $\frac{(10.13 \times 10^5) (5.667 \times 10^{-5})}{(8.314) (294.3)} = 2.346$

$$C_s \simeq 5$$
 Btu/lb-mol "R = 20.93 Metol K.

Thermodynamics And Its Applications 3rd Edition Solution Manual

Joseph Vincent Woodworth

Thermodynamics And Its Applications 3rd Edition Solution Manual:

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