

EXPERIMENT 2

ELECTROCHEMISTRY: ELECTROCHEMICAL CELL AND THERMODYNAMIC FUNCTIONS

Objectives

1. To construct electrochemical cells
2. To determine the net cell potentials for three electrochemical cells
3. To calculate the net cell potential based on the potentials of the half-reactions that occur and to compare your experimental and calculated values.
4. To measure the Gibbs energy based on the cell potentials.

Introduction

The use of electrochemical cells to convert the Gibbs energy stored in the constituent half-reactions into electrical work is of enormous industrial as well as fundamental significance. We have all used batteries and these are simply galvanic cells that are similar to those you will be constructing in this experiment. In the laboratory, a typical electrochemical cell has the following general construction:

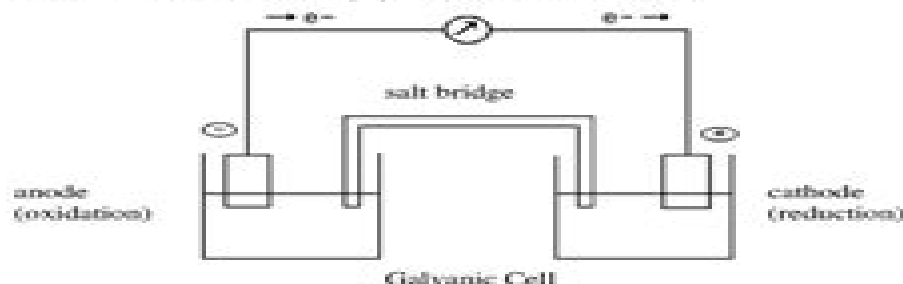


Figure 2.1: A schematic diagram of galvanic cell

The standard reduction potential is a quantitative measure of a substance's tendency to accept electrons under standard conditions (1 atm of gases and 1 M concentration) and is measured in volts. The more positive the reduction potential, the greater the tendency for reduction to occur. For the half-reaction that represents the oxidation of a metal, you can determine the standard oxidation potential (E°_{ox}) for the half-reaction. This value has the same numerical value as the standard reduction potential but the opposite sign. For example, aluminum ion has a standard reduction potential value of -1.65 V



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Table of Contents Thermodynamics Lab Manual Experiment

1. Understanding the eBook Thermodynamics Lab Manual Experiment
 - The Rise of Digital Reading Thermodynamics Lab Manual Experiment
 - Advantages of eBooks Over Traditional Books
2. Identifying Thermodynamics Lab Manual Experiment
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Thermodynamics Lab Manual Experiment
 - User-Friendly Interface
4. Exploring eBook Recommendations from Thermodynamics Lab Manual Experiment
 - Personalized Recommendations
 - Thermodynamics Lab Manual Experiment User Reviews and Ratings
 - Thermodynamics Lab Manual Experiment and Bestseller Lists
5. Accessing Thermodynamics Lab Manual Experiment Free and Paid eBooks
 - Thermodynamics Lab Manual Experiment Public Domain eBooks
 - Thermodynamics Lab Manual Experiment eBook Subscription Services
 - Thermodynamics Lab Manual Experiment Budget-Friendly Options
6. Navigating Thermodynamics Lab Manual Experiment eBook Formats
 - ePub, PDF, MOBI, and More
 - Thermodynamics Lab Manual Experiment Compatibility with Devices
 - Thermodynamics Lab Manual Experiment Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Thermodynamics Lab Manual Experiment
 - Highlighting and Note-Taking Thermodynamics Lab Manual Experiment
 - Interactive Elements Thermodynamics Lab Manual Experiment
8. Staying Engaged with Thermodynamics Lab Manual Experiment

-
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Thermodynamics Lab Manual Experiment
9. Balancing eBooks and Physical Books Thermodynamics Lab Manual Experiment
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Thermodynamics Lab Manual Experiment
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Thermodynamics Lab Manual Experiment
 - Setting Reading Goals Thermodynamics Lab Manual Experiment
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Thermodynamics Lab Manual Experiment
 - Fact-Checking eBook Content of Thermodynamics Lab Manual Experiment
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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