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Triumph 500cc Engine Rebuilding Manual

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web start with 8.3 mole H_2S and multiply it with 2 mole H_2O over 2 mole H_2S from the equation $8.3 \text{ mole } H_2S \times \frac{2 \text{ mole } H_2O}{2 \text{ mole } H_2S} \times 18 \text{ g } H_2O = 149.4 \text{ g } H_2O$ is your theoretical yield so percent yield $\frac{137.1 \text{ g } H_2O}{149.4 \text{ g } H_2O} \times 100 = 91.77\%$ problems 6 8 are similar to problem 5

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you have 0.216 moles $\times$ 32 grams/mole = 6.9 grams of O_2

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calculated mass of the product expected $\times 100$ in this reaction a yield of 80 is anticipated

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web a write a balanced equation for the reaction $\text{Al} + \text{O}_2 \rightarrow \text{Al}_2\text{O}_3$ b determine the theoretical yield of Al_2O_3 160.0

g $\text{Al} \times \frac{1 \text{ mol Al}}{27 \text{ g Al}} \times \frac{2 \text{ mol Al}_2\text{O}_3}{3 \text{ mol Al}} \times 101.96 \text{ g Al}_2\text{O}_3 = 60.0 \text{ g Al}_2\text{O}_3$ c determine the percent yield

yield actual $\times 100 = \frac{60.0 \text{ g}}{100 \text{ g}} \times 100 = 60.0\%$ theoretical 100.0 g

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web 7 Na₂C₂O₄ is the limiting reactant percent yield 86.6% only four molecules can be made 9 this amount cannot be

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C is the limiting reagent 4.33 g of H₂ are left over 13

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CH₃OH can be produced by the following reaction $2\text{H}_2 + \text{CO} \rightarrow \text{CH}_3\text{OH}$ calculate the theoretical yield of CH₃OH if 68.5 g of CO is

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