

KEY

Thermochem and Collision theory Study Guide

What is enthalpy? (define it and indicate its symbol)

heat content of a reaction or process. H

What is thermochemistry?

Study of heat (E) changes of a reaction

What is an exothermic reaction?

a reaction that gives off heat

Explain why an exothermic reaction have a negative enthalpy change?

b/c reactants lose heat to become products. (The overall ΔH is negative)

What is an endothermic reaction?

a reaction that absorbs heat

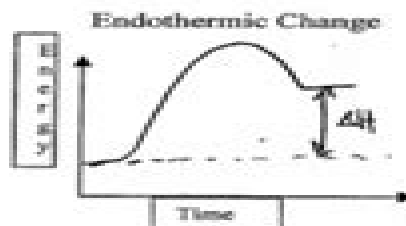
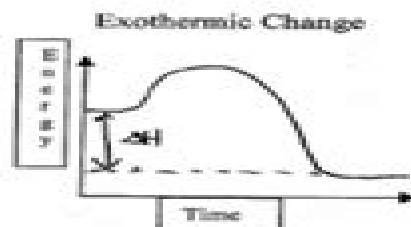
Explain why an endothermic reaction have a positive enthalpy change?

b/c reactants absorb heat to become products (The overall ΔH is +)

Complete the following chart:

Type of Reaction	Sign of ΔH_{rxn}	Which has more energy: reactants or products?	What temperature change would you see?
Exothermic	-	Reactants	Temp would $\uparrow$
Endothermic	+	Products	Temp. would $\downarrow$

Make enthalpy diagrams for the two chemical reactions from above, showing clearly the amount of energy released or gained.



Does the energy go from the surrounding to the chemicals or from the chemicals to the surroundings in an exothermic reaction? from chemicals to surroundings

Does the energy go from the surrounding to the chemicals or from the chemicals to the surroundings in an endothermic reaction? from surroundings to chemicals

Classify each of the following as an exothermic or endothermic process.

Melting ice cubes endothermic

Burning a candle exothermic

Evaporation of water endothermic

Baking Bread endothermic

Splitting a gas molecule apart endothermic (breaking bond!)

Formation of snow in clouds exothermic (freezing!)

What type of reactions are the following based on the enthalpy?

$C + H_2O \rightarrow CO + H_2$ $\Delta H_{rxn} = +113 \text{ kJ}$ endothermic

$2H_2O_2 \rightarrow 2H_2O + O_2$ $\Delta H = -190 \text{ kJ}$ exothermic

$6CO_2 + 6H_2O + 2803 \text{ kJ} \rightarrow C_6H_{12}O_6 + 6O_2$ endothermic

$3NO_2 + H_2O \rightarrow 2HNO_3 + NO + 138 \text{ kJ}$ exothermic

$P_4 + 5O_2 \rightarrow 3013 \text{ kJ} + P_4O_{10}$ exothermic

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**Frances M. Tappan, Patricia J.
Benjamin**



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Advanced Research on Material Engineering, Chemistry, Bioinformatics II Helen Zhang, David Jin, 2012-06-04 Selected peer reviewed papers from the 2012 2nd International Conference on Material Engineering Chemistry Bioinformatics MECB 2012 July 14 15 Xi an China

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