

Derive the following equations:

1. Show the following thermodynamic equations:

$$dC_p = \frac{C_p}{T} dT + \frac{C_p}{T^2} dT^2 + \frac{C_p}{T^3} dT^3 + \dots$$

(a) is this equation a differential or an integral thermodynamic equation?

$$dC_p = \frac{C_p}{T} dT + \frac{C_p}{T^2} dT^2 + \frac{C_p}{T^3} dT^3 + \dots$$

(b) what equation the integral of the above differential equation (C_p vs T) would be?

$$C_p(T) = \frac{C_p(T_0)}{T_0} + \frac{C_p(T_0)}{T_0^2} (T - T_0) + \frac{C_p(T_0)}{T_0^3} (T - T_0)^2 + \dots$$

(c) what equation the integral of the above differential equation (C_p vs T) would be?

$$C_p(T) = \frac{C_p(T_0)}{T_0} + \frac{C_p(T_0)}{T_0^2} (T - T_0) + \frac{C_p(T_0)}{T_0^3} (T - T_0)^2 + \dots$$

(d) what is the integral of the above differential equation (C_p vs T) would be?

$$C_p(T) = \frac{C_p(T_0)}{T_0} + \frac{C_p(T_0)}{T_0^2} (T - T_0) + \frac{C_p(T_0)}{T_0^3} (T - T_0)^2 + \dots$$

2. Derive the following thermodynamic equations: $dC_p = \frac{C_p}{T} dT + \frac{C_p}{T^2} dT^2 + \frac{C_p}{T^3} dT^3 + \dots$

and use the thermodynamic equations (which must include the appropriate sign with the appropriate sign) for the derivation of the following equations:

$$dC_p = \frac{C_p}{T} dT + \frac{C_p}{T^2} dT^2 + \frac{C_p}{T^3} dT^3 + \dots$$

3. If a solid substance (diamond) was compressed, would it be able to be used for the reaction: $dC_p = \frac{C_p}{T} dT + \frac{C_p}{T^2} dT^2 + \frac{C_p}{T^3} dT^3 + \dots$

$$dC_p = \frac{C_p}{T} dT + \frac{C_p}{T^2} dT^2 + \frac{C_p}{T^3} dT^3 + \dots$$

$$dC_p = \frac{C_p}{T} dT + \frac{C_p}{T^2} dT^2 + \frac{C_p}{T^3} dT^3 + \dots$$

4. If a solid substance (diamond) was compressed, would it be able to be used for the reaction: $dC_p = \frac{C_p}{T} dT + \frac{C_p}{T^2} dT^2 + \frac{C_p}{T^3} dT^3 + \dots$

$$dC_p = \frac{C_p}{T} dT + \frac{C_p}{T^2} dT^2 + \frac{C_p}{T^3} dT^3 + \dots$$

5. If a solid substance (diamond) was compressed, would it be able to be used for the reaction: $dC_p = \frac{C_p}{T} dT + \frac{C_p}{T^2} dT^2 + \frac{C_p}{T^3} dT^3 + \dots$

$$dC_p = \frac{C_p}{T} dT + \frac{C_p}{T^2} dT^2 + \frac{C_p}{T^3} dT^3 + \dots$$

6. If a solid substance (diamond) was compressed, would it be able to be used for the reaction: $dC_p = \frac{C_p}{T} dT + \frac{C_p}{T^2} dT^2 + \frac{C_p}{T^3} dT^3 + \dots$

a. Show a balanced equation for the reaction:



b. If a solid substance (diamond) was compressed, would it be able to be used for the reaction: $dC_p = \frac{C_p}{T} dT + \frac{C_p}{T^2} dT^2 + \frac{C_p}{T^3} dT^3 + \dots$

$$dC_p = \frac{C_p}{T} dT + \frac{C_p}{T^2} dT^2 + \frac{C_p}{T^3} dT^3 + \dots$$

c. If a solid substance (diamond) was compressed, would it be able to be used for the reaction: $dC_p = \frac{C_p}{T} dT + \frac{C_p}{T^2} dT^2 + \frac{C_p}{T^3} dT^3 + \dots$



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