

Marking Scheme

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Figure 1  
Section A

1. (a) (i)  $3\text{Ca}^{2+} + 2\left[\begin{array}{c} 18 \\ \text{N} \\ 10 \end{array}\right]^{3-}$  (03)
- (ii) None (00)  
Cations are positively charged (01)  
Cations are oppositely charged (01)
- (iii) (i) Subgroups are atoms of the same element with same proton number but different neutron numbers. (03)
- (ii)
- |                       | Number of protons | Number of neutrons |
|-----------------------|-------------------|--------------------|
| $^{40}_{20}\text{Ca}$ | 20                | 20                 |
- (03)
- (iii) (i) They have the same number of electrons in their outermost shells. (01)  
(ii) It has similar chemical properties to Ca does, then can replace some of the Ca required. (01)
2. (a) (i)  $2\text{AgCl} \rightarrow 2\text{Ag} + \text{Cl}_2$  (01)
- (ii) The molar mass of Ag decreases and the oxidation no. of Cl increases. (01)
- (iii) No. of moles of  $\text{AgCl} = \frac{1.36}{143.5} = 0.0095$  (01)  
No. of moles of Ag = 2 x No. of moles of  $\text{AgCl}$  (01)  
Mass of Ag that can be obtained =  $0.0095 \times 108 \times 2 = 2.07$  (01)  
= 2.04g (01)
- (b) (i) The black oxide changes to white brown color (01)  
(ii) The metal obtained can conduct electricity. (01)  
(iii)  $\text{CuO} + \text{H}_2 \rightarrow \text{Cu} + \text{H}_2\text{O}$  (01)  
(iv) Hydrogen is explosive + flammable. (01)  
(v) No. The reactivity of Cu and Ag can only be compared using the same reaction. (01)

3. (a) (i)  (03)
- (ii) (i) It is a solution of polymeric molecules of different chain lengths. (01)  
(ii) conductive (01)  
(iii) Electrically bonding metals / not both / not both should be used because terminal is flammable. (01)  
or, When the solution in a burner evaporated because polymer vapour is formed. (01)  
(iv) addition polymerisation (01)
- (iii) (i) Electrically bonding can be prevented. (01)  
(ii) Plastic does not conduct easily. (01)  
(iii) Low density (01)
3. Chemical Knowledge (03)
- Draw two of the metal strips into the beakers from a standard cell.  
Measure the voltage of the cell using the multimeter.  
Complete the circuit circuit.
- Control variables in the experiment such as:  
- the size of the strips  
- the separation/position between strips should be the same in each trial  
- the temperature
- The highest voltage can be obtained using a magnesium strip and a copper strip. (It is because among the three metals, Mg occupies the highest position in the e.m.f. and Cu the lowest position).
- Adjust the distance/position between the two metal strips until the maximum voltage (voltage) is obtained.
- Algebraic Calculation (03)

3. (a) (i) Sodium hydroxide is very corrosive. (01)  
(ii) The sodium hydroxide increase (01)
- (b) (i) Copper cannot displace H<sup>+</sup>(aq) from HCl(aq). (01)  
(ii) Add Zn/Mg/Pb to HCl(aq). (01)
- (c) (i) When water is added to concentrated  $\text{H}_2\text{SO}_4$ , a lot of heat is produced. The heat can cause splashing out of the concentrated acid solution. (01)  
(ii) Add concentrated  $\text{H}_2\text{SO}_4$  to water slowly and stir the mixture. (01)
4. (a) (i) Chemicals containing sulphuric acid are harmful to marine life/ aquatic organisms. (01)
- (ii) (i)  $\text{SO}_4^{2-} + \text{H}_2\text{O} \rightarrow \text{HSO}_4^- + \text{OH}^- + \text{H}^+$  (01)  
(ii)  $\text{Cu}_2\text{O}^{2+} + 2\text{OH}^- + \text{H}^+ \rightarrow 2\text{Cu}^{2+} + \text{H}_2\text{O}$  (01)  
(iii)  $\text{Cu}_2\text{O}^{2+} + \text{H}^+ + \text{HSO}_4^- \rightarrow 2\text{Cu}^{2+} + \text{HSO}_4^- + \text{H}_2\text{O}$  (01)
- (b)  $\text{MnO}_2 + \text{H}_2\text{C}_2\text{O}_4 + \text{H}_2\text{SO}_4 \rightarrow \text{CO}_2$  (01)
5. (a) Any TWO of the following (03)
- Substances and atoms can be represented by a semi-structural formula. The general formula for pentane and octane is  $\text{C}_5\text{H}_{12}$  and  $\text{C}_8\text{H}_{18}$  respectively differ by one  $-\text{CH}_2-$ .
  - There are gradual changes in physical properties among the members of a homologous series. The boiling point / melting point / viscosity / density of octane is higher than that of pentane.
  - Members of the same homologous series have similar chemical properties. Both pentane and octane can undergo substitution reactions with  $\text{Br}_2/\text{FeBr}_3$ .
- (b) Octane, it has a higher percentage of carbon by mass. Its chance to undergo incomplete combustion is give carbon is higher. (01)
- (c) Any TWO of the following (01)
- $\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_3$   
 $\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_3$   
 $\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_3$

- Section B
5. (a) (i)  $\frac{10}{100} = \frac{0.4}{x}$  (01)  
mole ratio (01)  
 $x = \frac{0.4 \times 100}{10} = 4$  (01)  
Empirical formula of X is  $\text{PbO}_4$  (01)
- (ii) Let mole ratio of  $\text{PbO}$  to  $\text{PbO}_2$  be  $x : y$  (01)  
 $\frac{0.4375 \times 207}{207} = \frac{0.5625 \times 239}{239} = \frac{x}{y}$  (01)  
 $x = \frac{0.4375 \times 207 \times y}{0.5625 \times 239} = \frac{1}{4}$  (01)  
X is a mixture of  $\text{PbO}$  and  $\text{PbO}_2$ , its mole ratio of 1 : 4. (01)  
or, X is not a mixture. In X, two-third of the total atoms are an oxidation number +2, while one-third is an oxidation number +4.
6. (a) (i) Hydrogen (01)  
(ii) manufacture of  $\text{NH}_3$  / manufacture of  $\text{CH}_3\text{NH}_2$  / manufacture of hydrochloric acid / fertilising of vegetable crop (01)
- (b) (i) Chlorine (01)  
(ii) At very low concentrations,  $\text{Cl}_2$  will be formed +  $\text{HCl}$  (aq) discharged. (01)  
Fraction of  $\text{OH}^-$  in  $\text{HCl}$  is higher than that of  $\text{Cl}^-$  /  $\text{OH}^-$  is a stronger reducing agent than  $\text{Cl}^-$  /  $\text{OH}^-$  have electron more readily than  $\text{Cl}^-$ . (01)
- (c) It is a reducing agent/ its from electrons (01)  
 $\text{MnO}_2$  is an oxidising agent /  $\text{MnO}_2$  gains electrons (01)  
When the cell is connected to an external circuit, electrons will flow through the external circuit. (01)

# Wjec Chemistry Marking Scheme 2005

**David Ballard, Rhodri Thomas**



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