

Thermochemistry Practice Problems ANSWER KEY

1. Calculate the standard enthalpy of reaction for the process



$$\Delta H = -176 \text{ kJ}$$

2. Calculate the heat evolved in the combustion of ethyl alcohol, ΔH_{comb}



$$\Delta H = -1366.7 \text{ kJ}$$

3. Calculate the heat required to decompose limestone, CaCO_3 , to lime, CaO , and carbon dioxide, with all substances at standard conditions.



$$\Delta H^\circ_f \text{ CaCO}_3(s) = -1206.9 \text{ kJ/mol}$$

$$\Delta H^\circ_f \text{ CaO}(s) = -635.1 \text{ kJ/mol}$$

$$\Delta H^\circ_f \text{ CO}_2(g) = -393.5 \text{ kJ/mol}$$

$$\Delta H = 178.3 \text{ kJ}$$

4. $\text{CH}_4(g) + 3 \text{Cl}_2(g) \rightarrow \text{CHCl}_3(l) + 3 \text{HCl}(g) \quad \Delta H = ?$



$$\Delta H = -305.2 \text{ kJ}$$

5. Cyanamide, CH_2N_2 , is a weak acid that is sometimes used as a fertilizer. Calculate the standard enthalpy of formation for cyanamide, given the following standard enthalpies of reaction:



$$\Delta H = 62.1 \text{ kJ}$$

6. The standard molar enthalpy change is -905.2 kJ for the oxidation of ammonia.



Calculate the standard molar enthalpy of formation for ammonia based on the following standard enthalpies of formation: $\Delta H^\circ_f \text{ NO}(g) = 90.3 \text{ kJ/mol}$ and $\Delta H^\circ_f \text{ H}_2\text{O}(g) = -241.8 \text{ kJ/mol}$

$$-46.1 \text{ kJ/mol}$$

7. The thermite reaction: $2 \text{Al}(s) + \text{Fe}_2\text{O}_3(s) \rightarrow \text{Al}_2\text{O}_3(s) + 2 \text{Fe}(s) + 852 \text{ kJ}$

If you begin with 10.0 g of Al and excess Fe_2O_3 , how many kJ of heat are evolved at constant pressure?

$$158 \text{ kJ}$$

Thermochemistry Practice Calculations Answers

Peter J. Mikulecky, Chris Hren



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Table of Contents Thermochemistry Practice Calculations Answers

1. Understanding the eBook Thermochemistry Practice Calculations Answers
 - The Rise of Digital Reading Thermochemistry Practice Calculations Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Thermochemistry Practice Calculations Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Thermochemistry Practice Calculations Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Thermochemistry Practice Calculations Answers
 - Personalized Recommendations

- Thermochemistry Practice Calculations Answers User Reviews and Ratings
- Thermochemistry Practice Calculations Answers and Bestseller Lists
- 5. Accessing Thermochemistry Practice Calculations Answers Free and Paid eBooks
 - Thermochemistry Practice Calculations Answers Public Domain eBooks
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- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Thermochemistry Practice Calculations Answers
 - Highlighting and Note-Taking Thermochemistry Practice Calculations Answers
 - Interactive Elements Thermochemistry Practice Calculations Answers
- 8. Staying Engaged with Thermochemistry Practice Calculations Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Thermochemistry Practice Calculations Answers
- 9. Balancing eBooks and Physical Books Thermochemistry Practice Calculations Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Thermochemistry Practice Calculations Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Thermochemistry Practice Calculations Answers
 - Setting Reading Goals Thermochemistry Practice Calculations Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Thermochemistry Practice Calculations Answers
 - Fact-Checking eBook Content of Thermochemistry Practice Calculations Answers

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
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