

**CHEMISTRY 161**  
**THERMO WORKSHEET**  
**FALL 2015**

1. A system absorbs 52.5 J of energy as it expands to a volume of 50.0 L. If the external pressure is 0.800 atm and the change in internal energy is -112 J, calculate the initial volume.

2. Use Hess's law to find the heat of combustion of methane,



using the following thermochemical equations.



3. How much glucose must be metabolized to completely evaporate 1.00 g of water? The heat of combustion of glucose is 2803 kJ/mole.
4. The standard enthalpy of formation of  $\text{NH}_3$  is -46.1 kJ/mole. What is  $\Delta H^\circ_{\text{rxn}}$  for the following reactions?
- a.  $\text{N}_2(g) + 3 \text{H}_2(g) \rightarrow 2\text{NH}_3(g)$       b.  $\text{NH}_3(g) \rightarrow \text{N}_2(g) + 3 \text{H}_2(g)$
5. The heat of combustion of pentane ( $\text{C}_5\text{H}_{12}$ ) is -3535 kJ/mole
- a. Find the heat given off when 1.00 kJ of pentane is burned.
- b. How many grams of pentane must be burned to increase the temperature of 1.00 kg of water from 20.0°C to 90.0°C?
6. Brass has a density of 8.40 g/cm<sup>3</sup> and a specific heat of 0.385 J/°C. A 15.2 cm<sup>3</sup> piece of brass at an initial temperature of 163°C is dropped into an insulated container with 150.0 g H<sub>2</sub>O at 22.4°C. What will be the final temperature of the brass-water mixture?

# Thermochemistry Calculations 1 Answers Keys

**Peter J. Mikulecky, Katherine  
Brutlag, Michelle Rose Gilman, Brian  
Peterson**



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## **Table of Contents Thermochemistry Calculations 1 Answers Keys**

1. Understanding the eBook Thermochemistry Calculations 1 Answers Keys
  - The Rise of Digital Reading Thermochemistry Calculations 1 Answers Keys
  - Advantages of eBooks Over Traditional Books
2. Identifying Thermochemistry Calculations 1 Answers Keys
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Thermochemistry Calculations 1 Answers Keys
  - User-Friendly Interface
4. Exploring eBook Recommendations from Thermochemistry Calculations 1 Answers Keys
  - Personalized Recommendations
  - Thermochemistry Calculations 1 Answers Keys User Reviews and Ratings

- Thermochemistry Calculations 1 Answers Keys and Bestseller Lists
- 5. Accessing Thermochemistry Calculations 1 Answers Keys Free and Paid eBooks
  - Thermochemistry Calculations 1 Answers Keys Public Domain eBooks
  - Thermochemistry Calculations 1 Answers Keys eBook Subscription Services
  - Thermochemistry Calculations 1 Answers Keys Budget-Friendly Options
- 6. Navigating Thermochemistry Calculations 1 Answers Keys eBook Formats
  - ePub, PDF, MOBI, and More
  - Thermochemistry Calculations 1 Answers Keys Compatibility with Devices
  - Thermochemistry Calculations 1 Answers Keys Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Thermochemistry Calculations 1 Answers Keys
  - Highlighting and Note-Taking Thermochemistry Calculations 1 Answers Keys
  - Interactive Elements Thermochemistry Calculations 1 Answers Keys
- 8. Staying Engaged with Thermochemistry Calculations 1 Answers Keys
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Thermochemistry Calculations 1 Answers Keys
- 9. Balancing eBooks and Physical Books Thermochemistry Calculations 1 Answers Keys
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Thermochemistry Calculations 1 Answers Keys
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Thermochemistry Calculations 1 Answers Keys
  - Setting Reading Goals Thermochemistry Calculations 1 Answers Keys
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Thermochemistry Calculations 1 Answers Keys
  - Fact-Checking eBook Content of Thermochemistry Calculations 1 Answers Keys
  - Distinguishing Credible Sources

13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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