

½CHEM 101 Worksheet, Chapter 6: Thermochemistry

Multiple choice questions 1-10 (page 1): circle the correct answer(A-E)

Show your work questions 11-13 (pages 2-3): give the complete answer in the space provided

R = 8.314 J/(mol.K) = 0.08206 L.atm/(mol.K) 1 cal = 4.184 J 0 °C = 273 K

1. Calculate the heat effect q (in kJ) when 112 g of iron (Fe(s)) is heated from 20 °C to 300 °C. The specific heat of Fe(s) = 0.48 J/g.K

- A. +6.53 kJ B. +29.7 kJ C. -15.5 kJ D. -29.7 kJ E. +15.5 kJ

2. The dissolution reaction for ammonium chloride: $\text{NH}_4\text{Cl(s)} \rightarrow \text{NH}_4^+(\text{aq}) + \text{Cl}^-(\text{aq})$ is endothermic. When ammonium chloride is dissolved in water in an insulated (adiabatic) system, the temperature will:

- A. increase B. decrease C. stay the same

3. Consider the reaction: $\text{CaCO}_3(\text{s}) \rightarrow \text{CaO(s)} + \text{CO}_2(\text{g})$. $\Delta H_{\text{rxn}} = +178 \text{ kJ}$ What is the enthalpy change when 20.0 kg CaO is formed?

- A. $+6.36 \times 10^4 \text{ kJ}$ B. $-6.36 \times 10^4 \text{ kJ}$ C. $+6.36 \times 10^4 \text{ J}$ D. 63.6 kJ E. +178 kJ

4. Consider the reaction: $2\text{NO} + \text{O}_2 \rightarrow 2\text{NO}_2$. $\Delta H_{\text{rxn}} = -114.6 \text{ kJ}$ As always, the ΔH is "for the reaction as written"! Calculate the enthalpy change ΔH , when 12.6 g NO_2 is produced.

- A. -31.4 kJ B. +31.4 kJ C. -15.7 kJ D. +15.7 kJ E. none of these

5. How much heat is associated with the reaction that produces 256.1 g of ammonia, NH_3 ?

$\text{N}_2(\text{g}) + 3 \text{H}_2(\text{g}) \rightarrow 2 \text{NH}_3(\text{g})$ $\Delta H_{\text{rxn}}^\circ = -92.45 \text{ kJ}$

- A. 92.45 kJ of heat is released B. 695.1 kJ of heat is released C. 184.9 kJ of heat is absorbed
D. 184.9 kJ of heat is released E. 92 kJ of heat is released

6. Which two of the following are **not** formation reactions?

1. $\frac{1}{2}\text{N}_2(\text{g}) + \frac{1}{2}\text{H}_2(\text{g}) \rightarrow \text{NH}_3(\text{g})$ 2. $\text{C(s)} + \text{O}_2(\text{g}) \rightarrow \text{CO}_2(\text{g})$ 3. $2\text{SO}_2(\text{g}) + \text{O}_2(\text{g}) \rightarrow 2\text{SO}_3(\text{g})$
4. $2\text{Fe(s)} + \frac{1}{2}\text{O}_2(\text{g}) \rightarrow \text{Fe}_2\text{O}_3(\text{s})$ 5. $\frac{1}{2}\text{H}_2(\text{g}) + \frac{1}{2}\text{I}_2(\text{g}) \rightarrow \text{HI(g)}$

- A. 1 and 2 B. 3 and 4 C. 2 and 3 D. 4 and 5 E. 3 and 5

7. Consider the thermochemical reaction equation: $2\text{CO(g)} + \text{O}_2(\text{g}) \rightarrow 2\text{CO}_2(\text{g})$ $\Delta H = -198 \text{ kJ}$

What is the enthalpy change for the reaction: $\text{CO}_2(\text{g}) \rightarrow \text{CO(g)} + \frac{1}{2}\text{O}_2(\text{g})$

- A. +198 kJ B. -99 kJ C. +99 kJ D. +396 kJ E. -198 kJ

8. Calculate the enthalpy change for the reaction: $2\text{KClO}_3 \rightarrow 2 \text{KCl} + \text{O}_2(\text{g})$

Standard enthalpy of formation values: $\Delta H_f^\circ(\text{KClO}_3) = -389 \text{ kJ/mol}$ $\Delta H_f^\circ(\text{KCl}) = -436 \text{ kJ/mol}$

- A. -94 kJ B. -47 kJ C. +94 kJ D. +47 kJ
E. Cannot be calculated because $\Delta H_f^\circ(\text{O}_2(\text{g}))$ is not given

9. Given the following two thermochemical equations:

1. $2\text{H}_2(\text{g}) + \text{O}_2(\text{g}) \rightarrow 2\text{H}_2\text{O(l)}$ $\Delta H = -572 \text{ kJ}$

2. $\text{H}_2(\text{g}) + \text{O}_2(\text{g}) \rightarrow \text{H}_2\text{O}_2(\text{l})$ $\Delta H = -188 \text{ kJ}$

What is the enthalpy change for the reaction: $2\text{H}_2\text{O}_2(\text{l}) \rightarrow 2\text{H}_2\text{O(l)} + \text{O}_2(\text{g})$?

- A. -98 kJ B. +98 kJ C. +196 kJ D. -196 kJ E. -670 kJ

10. Which of the following statements is correct for a combustion reaction in an adiabatic "bomb" calorimeter?

- A. the pressure is constant B. the temperature is constant C. the volume is constant
D. the calorimeter constant's C units are J/g.K E. $\Delta T(\text{K}) = \Delta T(^{\circ}\text{C}) + 273.15$

Show your work for the following questions.

Thermochemistry Multiple Choice

Sharma K.K. & Sharma L.K.



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A Textbook of Physical Chemistry, 6th Edition Sharma K.K. & Sharma L.K., 2016 A Textbook of Physical Chemistry **IIT Chemistry-II**,

Psychometrics R. Michael Furr, 2021-08-13 In this fully revised Fourth Edition of Psychometrics An Introduction author R Michael Furr centers his presentation around a conceptual understanding of psychometric core issues such as scales reliability and validity Focusing on purpose rather than procedure and the why rather than the how to this accessible book uses a wide variety of examples from behavioral science research so readers can see the importance of psychometric fundamentals in research By emphasizing concepts logic and practical applications over mathematical proofs this book gives students an appreciation of how measurement problems can be addressed and why it is important to address them The book offers readers the most contemporary views of topics in psychometrics available in the nontechnical psychometric literature

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