

Chemistry 30 Thermochemistry Quiz #2 Practice

Use the following information to answer the next question:

The combustion of ethanol ($\text{C}_2\text{H}_5\text{OH}$) at 298 K, standard state ($\text{C}_2\text{H}_5\text{OH}(l)$), can be represented by the following thermochemical equation:



Equation:



$$\Delta H^\circ = -1366.8 \text{ kJ}$$



$$\Delta H^\circ = -455.6 \text{ kJ}$$

4. The enthalpy change for the combustion of ethanol as written in the thermochemical equation is:

- a. -1366.8 kJ
- b. -455.6 kJ
- c. -1366.8 kJ
- d. -455.6 kJ

Use the following information to answer the next question:

When carbon is completely combusted (combustion), it represents the following equation:



5. The maximum heat enthalpy information with the reaction is

- a. -393.5 kJ/mol
- b. -393.5 kJ/mol
- c. -393.5 kJ/mol
- d. -393.5 kJ/mol

Thermochemistry Question Answer Guide

Robert K. Wismer



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Cooper, Thomas J. Greenbowe, 2005 For courses in Methods of Teaching Chemistry Useful for new professors chemical educators or students learning to teach chemistry Intended for anyone who teaches chemistry or is learning to teach it this book examines applications of learning theories presenting actual techniques and practices that respected professors have used to implement and achieve their goals Each chapter is written by a chemist who has expertise in the area and who has experience in applying those ideas in their classrooms This book is a part of the Prentice Hall Series in Educational Innovation for Chemistry Nature Of Chemistry Volume - 1 Aditya Tripathi, **General Chemistry Study Guide Sixth Edition** Darrell D. Ebbing, 1999 *Study Guide for General Chemistry and College Chemistry, Eighth Editions* by Holtzclaw and Robinson Norman E. Griswold, Henry Fuller Holtzclaw, 1988 **Active Learning Guide for Chemistry** Donald J. DeCoste, 2007 **Student Study Guide to Accompany Petrucci's General Chemistry** Robert K. Wismer, 1985

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