

Thermochemistry Worksheet #1

AP Chemistry, Chapter 6

ANSWERS

- 1) +1000 J, endothermic (q is positive)
- 2) -2267 J, work done by system on the surroundings
- 3) $\Delta E = +300$ J, $q = -200$ J exothermic
- 4) 0.5354 kJ, 0.1280 kcal
- 5) $q = +4184$ J; at constant P, so $\Delta H = q$ or $\Delta H = +4184$ J;
 $w = -30.4$ J; $\Delta E = 4154$ J
- 6) $q = +6500$ J, $w = -9620$ J, $\Delta E = -3120$ J
- 7) a. endothermic, 1.8 kJ = 0.43 kcal
b. exothermic, -2.4 kcal = -10 kJ
c. exothermic, -0.025 kcal = -0.10 kJ
d. exothermic, -95.3 kJ = -22.8 kcal
- 8) 0.0329 J/g°C or 0.00786 cal/g°C
- 9) 59.7 kJ
- 10) -49.4 kJ/mol, exothermic
- 11) -289 kJ/mol H_2
- 12) 18.9 L
- 13) -5670 kJ/mol sucrose, -11200 kJ/mol stearic acid
- 14) $\Delta H_{rxn} = -822.3$ kJ
- 15) $\Delta H_{rxn} = -78.7$ kJ

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Peter J. Mikulecky, Chris Hren



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