REVIEW FOR THERMOCHEMISTRY TEST

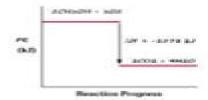
1. The combustion of methanol is shown by the following equation:

$$2CH_3OH_{(R)} + 3O_{2(R)} \longrightarrow 2CO_{2(R)} + 4H_2O_{(R)}$$

a) Using the data below, find the heat of reaction for the combustion of methanol.

$$C_{03} \cdot O_{21g3} \rightarrow CO_{21g3}$$
 $\Delta H = -393 \text{ kJ (x2)}$
 $2H_{21g3} \cdot O_{21g3} \rightarrow 2H_2O_{1g3}$ $\Delta H = -484 \text{ kJ (x2)}$
 $2C_{03} \cdot 4H_{21g3} \cdot O_{21g3} \rightarrow 2CH_2OH_{23}$ $\Delta H = -476 \text{ kJ (x (-1))}$
 $2C_{03} \cdot 2O_{21g3} \rightarrow 2CO_{21g3}$ $\Delta H = -786 \text{ kJ}$
 $4H_{21g3} \cdot 2O_{21g3} \rightarrow 4H_2O_{1g3}$ $\Delta H = -968 \text{ kJ}$
 $2CH_2OH_{23} \rightarrow 2CO_{21g3} \rightarrow 4H_2O_{1g3}$ $\Delta H = 476 \text{ kJ}$
 $2CH_2OH_{23} \rightarrow 2CO_{21g3} \rightarrow 2CO_{21g3} \cdot 4H_2O_{1g3}$ $\Delta H = 476 \text{ kJ}$
 $2CH_2OH_{23} \rightarrow 3O_{21g3} \rightarrow 2CO_{21g3} \cdot 4H_2O_{1g3}$ $\Delta H = -1278 \text{ kJ}$

- b) What is the molor enthalpy of combustion for methanol? AH = -1278 kJ/2 mol = -639 kJ/mol
- c) Express the energy changes for the target reaction in the form ΔH_e = ΔH_{enthalton} CH₂OH = -639 kJ/mol
- d) Draw the potential energy diagram to illustrate the energy changes for the target reaction.



Thermochemistry Test

Karl K. Irikura

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Table of Contents Thermochemistry Test

- 1. Understanding the eBook Thermochemistry Test
 - The Rise of Digital Reading Thermochemistry Test
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Thermochemistry Test
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Thermochemistry Test
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Thermochemistry Test
 - Personalized Recommendations
 - $\circ\,$ Thermochemistry Test User Reviews and Ratings
 - Thermochemistry Test and Bestseller Lists
- 5. Accessing Thermochemistry Test Free and Paid eBooks
 - Thermochemistry Test Public Domain eBooks
 - Thermochemistry Test eBook Subscription Services
 - Thermochemistry Test Budget-Friendly Options

- 6. Navigating Thermochemistry Test eBook Formats
 - o ePub, PDF, MOBI, and More
 - Thermochemistry Test Compatibility with Devices
 - Thermochemistry Test Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Thermochemistry Test
 - Highlighting and Note-Taking Thermochemistry Test
 - Interactive Elements Thermochemistry Test
- 8. Staying Engaged with Thermochemistry Test
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Thermochemistry Test
- 9. Balancing eBooks and Physical Books Thermochemistry Test
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Thermochemistry Test
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Thermochemistry Test
 - Setting Reading Goals Thermochemistry Test
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Thermochemistry Test
 - Fact-Checking eBook Content of Thermochemistry Test
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements

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