

## Acid or Base in Water Solution

- Acids produce  $\text{H}^+$  ions in presence of water.
- $\text{H}^+$  ions cannot exist alone, they exist as  $\text{H}_3\text{O}^+$  (hydronium ions).



Bases when dissolved in water gives  $\text{OH}^-$  ions.



- Bases soluble in water are called alkali.
- While diluting acids, it is recommended that the acid should be added to water and not water to acid because the process of dissolving an acid or a base in water is highly exothermic.



If water is added to acid, the heat generated may cause the mixture to splash out and cause burns and the glass container may also break due to excessive local heating.

**Adding water to acid may**

Cause mixture to splash out

Break the glass container

Mixing an acid or a base with  $\text{H}_2\text{O}$  results in decrease of concentration of ions ( $\text{H}_3\text{O}^+/\text{OH}^-$ ) per unit volume. Such a process is called as dilution.

## Strength of Acid and Base

Strength of acid or base can be estimated using universal indicator.

# Topic 10 Acids Bases And Salts Practice Questions

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