Covalent Bonding Worksheet

Covalent bonding occurs when two or more NON_METALS share electrons, attempting to attain a stable octet (8 outer electrons) in their outer shell for at least part of the time. Draw a Lewis dot diagram for each element listed. Circle the unpaired electrons that will be shared between the elements.

1.) H_{2 hydrogen} is diatomic

H + H → H → H-H Single Bond

2.) F₂ theoring in dianomic

 $F + F \rightarrow F \rightarrow F$

3.) Oz oxygen is dianomic

 $O + O \rightarrow$

 \rightarrow O=O

Double Bond

4.) No nitrogen is dianomic. Is this a triple bond?

N + N

5.) BF3 you need 3 fluoring atoms here

B + F

6.) Ammonia NH3 hist: how many hydrogen atoms are needed?

N + H

7.) Carbon dioxide CO2

O + C + O

8.) Methane CH4 careful here 4 hydrogen atoms needed

C + H

9.) Dihydrogen monoxide: the most dangerous substance on the planet. It has killed more people than any other substance known to mankind!!

H + O + H

10.) SO2 hint: one pair of electrons from sulfur must be slit up for this one to work.

O + S + O

Theories Of Covalent Bonding Practice Test

Arshad Iqbal

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