

### 3.3 Thermodynamics: Study Guide

Use this [outline](#) to review the key vocabulary for this unit.

1. Match the vocabulary words on the left with the appropriate definitions and units on the right.

|                   |                                                                                             |
|-------------------|---------------------------------------------------------------------------------------------|
| a. Thermal Energy | <u>A</u> Kinetic energy in transit from one object to another due to temperature difference |
| b. Temperature    | <u>B</u> Degree Celsius                                                                     |
|                   | <u>C</u> Joules                                                                             |
|                   | <u>D</u> The average kinetic energy of particles in an object                               |

2. In the table below, write in the appropriate formulas.

| Energy Transfer         | Rate of Energy Transfer                                                                                     | Thermal Conductivity        | Coefficient of Heat Conductivity |
|-------------------------|-------------------------------------------------------------------------------------------------------------|-----------------------------|----------------------------------|
| $Q = mc \cdot \Delta T$ | $P = \frac{c \Delta T}{L}$ or $P = \frac{Q}{t}$<br><u>Thermal conductance</u> $P = \sigma A c (\Delta T)^2$ | $K = \frac{PL}{A \Delta T}$ | $U = \frac{P}{A \cdot \Delta T}$ |

3. For each example, write the correct number for the law of thermodynamics it describes (0, 1, 2).
- 2 The law of entropy
  - 1 Thermal energy can only be increased in a system by adding heat from an external source or performing work within the system
  - 0 If two systems are separately found to be in thermal equilibrium with a third, then the two systems are in thermal equilibrium with each other
  - 2 Thermal energy flows only from warmer to cooler
  - 1 The law of conservation of energy applied to thermal systems
4. How much thermal energy would it take to raise the temperature of 1.1 (1 kg) of water by 10°C? The specific heat capacity of water is  $4184 \frac{J}{kg \cdot ^\circ C}$

$$Q = mc \Delta T$$

$$= 1.1 kg \cdot 4184 \frac{J}{kg \cdot ^\circ C} \cdot 10^\circ C$$

$$Q = 41,840 J$$

$$Q = \underline{41,840 J}$$

5. A materials scientist is trying to determine the specific heat capacity of an experimental material. She applies 30 kJ of thermal energy to a 5 kg sample of the material and records a 24°C temperature increase in the sample. What is the specific heat capacity of the material?

$$Q = mc \Delta T$$

$$c = \frac{Q}{m \Delta T}$$

$$c = 2500 \frac{J}{kg \cdot ^\circ C}$$

$$c = \underline{2500 \frac{J}{kg \cdot ^\circ C}}$$

$$= \frac{30000 J}{5 kg \cdot 24^\circ C}$$

# Thermodynamics Chemistry Study Guide Answers

**Arshad Iqbal**



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