

WHMIS Symbols

The shapes of the symbols used on labels have been chosen to show the nature of the hazard they represent.

Class A	COMPRESSED GAS • It is a gas kept under pressure. • Heat may cause the container to <u>explode</u>. • A drop or impact may cause the container to explode. E.g. steel cylinders of acetylene, oxygen, hydrogen, helium, chlorine, nitrogen, neon, argon etc. and fire extinguishers. • Compressed gases can be hazardous simply because they are under high pressure, and the gas itself can also be hazardous (like chlorine gas). If the gas itself is hazardous, it will have other appropriate hazard symbols along with the compressed gas symbol. • Handle with care, do not drop. • Keep away from heat or potential sources of ignition. Store in a designated area. Large cylinders must be properly secured with a chain.	
Class B	FLAMMABLE AND COMBUSTIBLE MATERIALS • The material is a potential fire hazard. It may burn at relatively low temperature. Sparks, flame or friction could ignite it. • May burst into flame spontaneously in air or release a flammable gas on contact with water. • Keep any of these materials away from heat sources and other combustible materials. Never smoke when working with or near the materials. Store in a cool, fire-proof area. Division 1: Flammable Gases: E.g. hydrogen, methane, propane. Division 2: Flammable Liquids: Flashpoint less than 37°C. These liquids catch on fire easily and have highly flammable fumes. E.g. gasoline, ethanol, methanol, diethyl ether. Division 3: Combustible Liquids: Flashpoint > 37°C E.g. diesel fuel, kerosene. These are less easily ignited than flammable liquids. Division 4: Flammable Solids: E.g. magnesium, sodium, beryllium. Division 5: Flammable Aerosols: E.g. most aerosol cans contain flammable propellants, also butane, propane in aerosol containers. Division 6: Flammable Reactive Materials: Materials that could spontaneously ignite in air (celluloid, lithium aluminum hydride) or in water (sodium).	

Whmis Student Handout

Richard Bailey

A red circular graphic with a gradient, appearing as a stylized arrow or a partial circle, located to the right of the author's name.

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 $\cos^2 x \cos x = \cos^2 x \sin^2 x \cos x = \cos^2 x (1 - \cos^2 x) \cos x = \cos^2 x \cos x - \cos^2 x \cos^3 x = \cos^2 x \cos x (1 - \cos x)$
let $y = \cos x$ then $\cos^2 x \cos x = 2\cos^2 x \cos x - 1 = 0$
 $2y^2 - y - 1 = 0$
 $y = \frac{1 \pm \sqrt{1 + 8}}{4} = \frac{1 \pm 3}{4}$
 $y = 1$ or $y = -\frac{1}{2}$

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a 2.00 L of 18.5 M H₂SO₄ concentrated sulfuric acid
b 100.0 mL of 3.8 × 10⁻⁵ M NaCN the minimum lethal concentration of sodium cyanide in blood serum

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