- Which quantities must be conserved in all chemical reactions?
  - A) mass, charge, density
  - (B) mass, charge, energy
  - C) charge, volume, density
  - D) charge, volume, energy
- 2 Given the balanced equation:

I + I → I2 Which statement describes the process represented by this equation?

- A) A bond is formed as energy is absorbed.
- (B) A bond is formed and energy is released.
- C) A bond is broken as energy is absorbed.
- D) A bond is broken and energy is released.
- Given the formula for hydrazine:

How many pairs of electrons are shared between the two nitrogen atoms?

- B) 2
- C) 3
- D) 4
- What is the number of pairs of electrons shared in a. molecule of No?
  - AT I
- B) 2
- D) 6
- 5. Which atom in the ground state has a stable valence electron configuration?
- B) Al
- C) Si
- D) Na
- An atom of which element has the strongest attraction for electrons in a chemical bond?
- A chlorine

- B) carbon
- C) phosphorus
- D) sulfur
- 7. Which bond is least polar?
  - A) As-Cl

B) Bi-Cl

C) P-C1

Given the electron dot diagram:

#### H:F:

The electrons in the bond between hydrogen and fluorine are more strongly attracted to the atom of

- A) hydrogen, which has the higher electronegativity
- (B) fluorine, which has the higher electronegativity
- C) hydrogen, which has the lower electronegativity
- fluorine, which has the lower electronegativity
- Given the reaction:

#### $M + 2 \text{ H2O} \rightarrow M(\text{OH})_2 + \text{H2}$

The metal represented by M is most likely a metal. from Group

- A) I
- C) 11
- D) 13
- Which type of substance can conduct electricity in the liquid phase but not in the solid phase?
  - (A) ionic compound
  - B) molecular compound
  - C) metallic element
  - D) nonmetallic element
- The chemical bonding in sodium phosphate, NasPO<sub>4</sub> , is classified as
  - A) ionic, only
  - B) metallic, only
  - Doth covalent and ionic
  - D) both covalent and metallic
- 12. Which type of bonding is found in all molecular substances?
  - (A) covalent bonding
- B) hydrogen bonding
  - C) ionic bonding
- D) metallic bonding
- Which type of substance is soft, has a low melting. point, and is a poor conductor of heat and electricity?
  - A) network solid
- - molecular solid D) ionic solid
- C) metallic solid

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