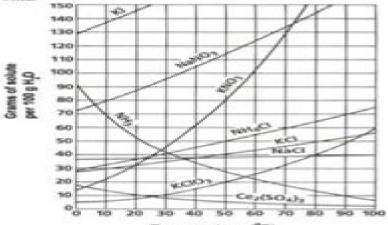
ame:	 D_{θ}

Solubility Curve Practice Problems

Directions: Find the mass of solute will dissolve in 100mL of water at the following temperatures?

- KNO₃ at 70°C = ______
- NaCl at 100°C=______
- NH_iCl at 90°C=
- Which of the above three substances is most soluble in water at 15°C. =
- Which of the above three substances is lest soluble in water at 15°C. =



Period:

	a description and an a contract of the party	
2. Solution	Saturated, super saturated or Unsaturated?	If unsaturated: How much more solute can dissolve in the solution?
a solution that contains 70g of NaNO3 at 30°C (in 100 mL H2O)		
a solution that contains 50g of NH ₄ Cl at 50°C (in 100 mL H ₂ O)		
a solution that contains 20g of KClO ₃ at 50°C (in 50 mL H ₂ O)		
a solution that contains 70g of KI at 0°C (in 200 mL H ₂ O)		

- 3. a. How many grams of KCl can dissolve at 5°C?
 - b. How many grams of KCl can dissolve at 25°C?
 - c. How many grams of Ce₂(SO₄)₃ can dissolve at 10°C?
 - d. How many grams of Ce₂(SO₄)₃ can dissolve at 50°C?
- 4. At 90°C, you dissolved 10 g of KCl in 100, g of water. Is this solution saturated or unsaturated?
- 5. A mass of 100 g of NaNO3 is dissolved in 100 g of water at 80°C.
 - a) Is the solution saturated or unsaturated?
 - b) As the solution is cooled, at what temperature should solid first appear in the solution? Explain.
- 6. Use the graph to answer the following two questions:

Which compound is most soluble at 20 °C?

Which is the least soluble at 40 °C?

Which substance on the graph is least soluble at 10°C?

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Table of Contents Unit 12 Solutions Solubility Curves Answer Key

- 1. Understanding the eBook Unit 12 Solutions Solubility Curves Answer Key
 - The Rise of Digital Reading Unit 12 Solutions Solubility Curves Answer Key
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Unit 12 Solutions Solubility Curves Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Unit 12 Solutions Solubility Curves Answer Key
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Unit 12 Solutions Solubility Curves Answer Key
 - Personalized Recommendations
 - Unit 12 Solutions Solubility Curves Answer Key User Reviews and Ratings
 - Unit 12 Solutions Solubility Curves Answer Key and Bestseller Lists

- 5. Accessing Unit 12 Solutions Solubility Curves Answer Key Free and Paid eBooks
 - Unit 12 Solutions Solubility Curves Answer Key Public Domain eBooks
 - Unit 12 Solutions Solubility Curves Answer Key eBook Subscription Services
 - Unit 12 Solutions Solubility Curves Answer Key Budget-Friendly Options
- 6. Navigating Unit 12 Solutions Solubility Curves Answer Key eBook Formats
 - ∘ ePub, PDF, MOBI, and More
 - Unit 12 Solutions Solubility Curves Answer Key Compatibility with Devices
 - Unit 12 Solutions Solubility Curves Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Unit 12 Solutions Solubility Curves Answer Key
 - Highlighting and Note-Taking Unit 12 Solutions Solubility Curves Answer Key
 - Interactive Elements Unit 12 Solutions Solubility Curves Answer Key
- 8. Staying Engaged with Unit 12 Solutions Solubility Curves Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Unit 12 Solutions Solubility Curves Answer Key
- 9. Balancing eBooks and Physical Books Unit 12 Solutions Solubility Curves Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Unit 12 Solutions Solubility Curves Answer Key
- 10. Overcoming Reading Challenges
 - $\circ\,$ Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Unit 12 Solutions Solubility Curves Answer Key
 - Setting Reading Goals Unit 12 Solutions Solubility Curves Answer Key
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Unit 12 Solutions Solubility Curves Answer Key
 - Fact-Checking eBook Content of Unit 12 Solutions Solubility Curves Answer Key
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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