# Specific Heat Worksheet

# $C_p = q/m\Delta T$ , where q = heat energy, m = mass, and T = temperature

 A 15.75-g piece of iron absorbs 1086.75 joules of heat energy, and its temperature changes from 25°C to 175°C. Calculate the heat capacity of iron, Cp = 46 J/92

1086.75 = 15.75g . 150 . Cp 1086.75= 2362.5 Cp

- How many joules of heat are needed to raise the temperature of 10.0 g of aluminum from 22°C to 55°C, if the specific heat of aluminum is 0.90 J/g°C? 9 = Cp.m.AT .90(10)(33) = 297 J
- To what temperature will a 50.0 g piece of glass raise if it absorbs 5275 joules of heat and its heat capacity is 0.50 J/g°C? The initial temperature of the glass is 20.0°C.

5275 = - 50( 50) AT Te = 120 231% AT = QII

 Calculate the heat capacity of a piece of wood if 1500.0 g of the wood absorbs 6.75×10<sup>4</sup> joules of heat, and its temperature changes from 32°C to 57°C.

67500 = Cp. 1500 - 25°C 1.8 J/gc

 100.0 mL of 4°C water is heated until its temperature is 37°C. If the specific heat of water is 4.18 J/g\*C, calculate the amount of heat needed to cause this rise in temperature.

C = 100g - 4.18 - 33°C 137947 13794

25.0 g of mercury is heated from 25°C to 155°C, and absorbs 455 joules of heat in the process. Calculate the heat capacity of mercury.

455 = Cp · 25g · 130 455 = 3258 Cp 1-14 = Cp

What is the Specific heat capacity of silver metal if 55.00 g of the metal absorbs 198 joules of heat and the temperature rises 15.0°C?

198 = Cp · 55g · 15°C Cp= . 24 J/gc

8. What is the change in temperature of 150.0 g of chloroform if it absorbs 1000.0 joules of heat, and the specific heat of chloroform is 0.96 J/g°C?

1000 = .96 (150g) AT 6.94 = AT

#### Answers

- 1. Co = .46 3/g°C
  - $2. \quad Q = 297 \text{ J}$
  - 3. Terral = 231 °C
  - C<sub>o</sub> =1.8 J/g\*C
  - 5. Q = 13800 ¥ 13794
  - 6. Co = 14 J/g\*C 7. Co = 24 J/g\*C
  - 8. AT = 6.9 °C

# **Using Specific Heat Answer Key**

**Robert Llewellyn Taylor** 

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