

**CHEMISTRY 161
THERMO WORKSHEET
FALL 2015**

1. A system absorbs 52.5 J of energy as it expands to a volume of 50.0 L. If the external pressure is 0.800 atm and the change in internal energy is -112 J, calculate the initial volume.

2. Use Hess's law to find the heat of combustion of methane,



using the following thermochemical equations.



3. How much glucose must be metabolized to completely evaporate 1.00 g of water? The heat of combustion of glucose is 2803 kJ/mole.
4. The standard enthalpy of formation of NH_3 is -46.1 kJ/mole. What is $\Delta H^\circ_{\text{rxn}}$ for the following reactions?
- a. $\text{N}_2(g) + 3 \text{H}_2(g) \rightarrow 2\text{NH}_3(g)$ b. $\text{NH}_3(g) \rightarrow \text{N}_2(g) + 3 \text{H}_2(g)$
5. The heat of combustion of pentane (C_5H_{12}) is -3535 kJ/mole
- a. Find the heat given off when 1.00 kJ of pentane is burned.
- b. How many grams of pentane must be burned to increase the temperature of 1.00 kg of water from 20.0°C to 90.0°C?
6. Brass has a density of 8.40 g/cm³ and a specific heat of 0.385 J/°C. A 15.2 cm³ piece of brass at an initial temperature of 163°C is dropped into an insulated container with 150.0 g H₂O at 22.4°C. What will be the final temperature of the brass-water mixture?

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SA Dillow



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