

Table F19 THERMODYNAMICS FORMULAE

Density (ρ)	$\rho = \frac{\text{mass}}{\text{volume}} = \frac{m}{V}$	Virial Equations $P = RT/(v - B)$ $B_0 = BP_0/RT_c$ $z = 1 + (B_0 P_0/T_c)$ $B^{(0)} = 0.083 - 0.422/T_r^{1.6}$ $B^{(1)} = 0.139 - 0.172/T_r^{4.2}$ $B_0 = B^{(0)} + \omega B^{(1)}$
Specific volume (v)	$v = \frac{\text{volume}}{\text{mass}} = \frac{V}{m} = \frac{1}{\rho}$	
Specific weight (γ)	$\gamma = \frac{\text{weight}}{\text{volume}} = \frac{mg}{V} = \rho g$	
Continuity equation	$m = \sum \rho VA = \sum pVA$	
Ideal-gas Equation of State	$pv = \frac{\bar{R}}{M} T$ $\bar{R}$ = universal gas constant M = molar mass of substance	
Energy balance for an open system	$Q - W = m_2(u_2 + p_2 v_2 + \frac{V_2^2}{2} + gz_2) - m_1(u_1 + p_1 v_1 + \frac{V_1^2}{2} + gz_1) \quad (\text{kJ})$ $Q - W = \Delta u + \Delta p v + \Delta \text{KE} + \Delta \text{PE}$	
Work of steady-flow, open system	$W_{\text{sh}} = \int v dp + \Delta \text{KE} + \Delta \text{PE}$	
Thermal efficiency of heat engine	$\eta = \frac{\dot{Q}_{\text{net}}}{\dot{Q}_{\text{in}}} = \frac{Q_{\text{in}} - Q_{\text{out}}}{Q_{\text{in}}}$ $\eta = \frac{T_H - T_L}{T_H}$	
Coefficient of performance (refrigerator)	$\text{COP}_R = \frac{\dot{Q}_{\text{out}}}{\dot{W}_{\text{in}}}$ $\text{COP}_R = \frac{T_L}{T_H - T_L}$	
Coefficient of performance (heat pump)	$\text{COP}_{\text{HP}} = \frac{\dot{Q}_{\text{out}}}{\dot{W}_{\text{in}}}$ $\text{COP}_{\text{HP}} = \frac{T_H}{T_H - T_L}$	
Change in entropy	$\Delta S = \int \left(\frac{\delta Q}{T} \right)_{\text{rev}} = S_2 - S_1 \quad (\text{kJ/K})$	
Change in entropy for open system	$\Delta S_{\text{sh}} = \left(\frac{\delta Q}{T} \right)_{\text{rev}} + s_1 \dot{m}_1 - s_2 \dot{m}_2 + \dot{S}_{\text{gen}} \quad (\text{kJ/K})$ $\frac{dS_{\text{sh}}}{dt} = \dot{\hat{Q}}_{\text{rev}} + \dot{m}_1 s_1 - \dot{m}_2 s_2 + \dot{S}_{\text{gen}} \quad (\text{kW/K})$	
Change in entropy for ideal gas	$dS = \int_1^2 c_p \frac{dT}{T} - R \ln \frac{P_2}{P_1}; \quad dS = \int_1^2 c_v \frac{dT}{T} + R \ln \frac{P_2}{P_1}; \quad \int_1^2 c_p \frac{dT}{T} = S_2 - S_1$	
Isentropic Turbine Efficiency	$\eta_T = \frac{\text{work of actual expansion}}{\text{work of reversible expansion}} = \frac{h_1 - h_{2s}}{h_1 - h_{2a}}$	
Adiabatic Compressor Efficiency	$\eta_C = \frac{\text{work of reversible compression}}{\text{work of actual compression}} = \frac{h_{2s} - h_1}{h_{2a} - h_1}$	

Expression for a
Reversible Process
(Expansion/
Compression)
in which
 $PV^\gamma = \text{constant}$

	Isothermal	Adiabatic, c.p.c. (γ)
ΔU	0	$\frac{nR}{\gamma-1}(T_2 - T_1)$ or $\frac{1}{\gamma-1}(P_2 V_2 - P_1 V_1)$
Q	$-nRT \ln\left(\frac{P_2}{P_1}\right)$	0
W	$nRT \ln\left(\frac{P_2}{P_1}\right)$	$\frac{nR}{\gamma-1}(T_2 - T_1)$ or $\frac{1}{\gamma-1}(P_2 V_2 - P_1 V_1)$
γ	1	c_p/c_v

Thermodynamics Property Tables Let Ta

Frederick Dominic Rossini



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John Joseph Flather,1915 **Plutonium Handbook** Oswald J. Wick,1967 **Thermophysical Properties of Gaseous**
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