

3.3 Thermodynamics: Study Guide

Use this [outline](#) to review the key vocabulary for this unit.

1. Match the vocabulary words on the left with the appropriate definitions and units on the right.

a. Thermal Energy	<u>A</u> Kinetic energy in transit from one object to another due to temperature difference
b. Temperature	<u>B</u> Degree Celsius
	<u>C</u> Joules
	<u>D</u> The average kinetic energy of particles in an object.

2. In the table below, write in the appropriate formulas.

Energy Transfer	Rate of Energy Transfer	Thermal Conductivity	Coefficient of Heat Conductivity
$Q = mc \cdot \Delta T$	$P = \frac{cA\Delta T}{L}$ or $P = \frac{Q}{t}$ <u>Thermal conductance</u> $P = \sigma A_c (T_1^4 - T_2^4)$	$k = \frac{PL}{A\Delta T}$	$U = \frac{P}{A \cdot \Delta T}$

3. For each example, write the correct number for the law of thermodynamics it describes (0, 1, 2).
- 2 The law of entropy
 - 1 Thermal energy can only be increased in a system by adding heat from an external source or performing work within the system
 - 0 If two systems are separately found to be in thermal equilibrium with a third, then the two systems are in thermal equilibrium with each other.
 - 2 Thermal energy flows only from warmer to cooler
 - 1 The law of conservation of energy applied to thermal systems
4. How much thermal energy would it take to raise the temperature of 1.1 (1 kg) of water by 10°C? The specific heat capacity of water is $4184 \frac{J}{kg \cdot ^\circ C}$

$$Q = mc \Delta T$$

$$= 1.1 kg \cdot 4184 \frac{J}{kg \cdot ^\circ C} \cdot 10^\circ C$$

$$Q = 41,840 J$$

$$Q = \underline{41,840 J}$$

5. A materials scientist is trying to determine the specific heat capacity of an experimental material. She applies 30 kJ of thermal energy to a 5 kg sample of the material and records a 24°C temperature increase in the sample. What is the specific heat capacity of the material?

$$Q = mc \Delta T$$

$$c = \frac{Q}{m \Delta T}$$

$$c = 2500 \frac{J}{kg \cdot ^\circ C}$$

$$c = \underline{2500 \frac{J}{kg \cdot ^\circ C}}$$

$$= \frac{30000 J}{5 kg \cdot 24^\circ C}$$

Thermodynamics Chemistry Study Guide Answers

Brendan G. Carr



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Study Guide James E. Brady, John R. Holum, 1993-02-11 Offers accurate lucid and interesting explanations of basic concepts and facts of chemistry while helping students develop skills in analytical thinking and problem solving Students are taught in a variety of ways to think of skills as tools that can be used to solve complex problems Several aids are included to help focus and inspire student interest frequent reference to common chemicals in commercial products numerous photographs of reactions in chapter practice exercises following worked examples

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Fundamental Concepts of Applied Chemistry Jayashree Ghosh, 2010-12 During the past few decades the growth of applied chemistry has been phenomenal and its applications have an expansive field including Chemical and Medico Biological disciplines I take pleasure in presenting the book Fundamental concepts of applied chemistry The book is published to provide a concise text book that encompasses important branches like pharmaceutical Biological polymer leather and Agricultural Chemistry

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Table of Contents Thermodynamics Chemistry Study Guide Answers

1. Understanding the eBook Thermodynamics Chemistry Study Guide Answers
 - The Rise of Digital Reading Thermodynamics Chemistry Study Guide Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Thermodynamics Chemistry Study Guide Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Thermodynamics Chemistry Study Guide Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Thermodynamics Chemistry Study Guide Answers
 - Personalized Recommendations
 - Thermodynamics Chemistry Study Guide Answers User Reviews and Ratings
 - Thermodynamics Chemistry Study Guide Answers and Bestseller Lists
5. Accessing Thermodynamics Chemistry Study Guide Answers Free and Paid eBooks
 - Thermodynamics Chemistry Study Guide Answers Public Domain eBooks
 - Thermodynamics Chemistry Study Guide Answers eBook Subscription Services
 - Thermodynamics Chemistry Study Guide Answers Budget-Friendly Options

6. Navigating Thermodynamics Chemistry Study Guide Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Thermodynamics Chemistry Study Guide Answers Compatibility with Devices
 - Thermodynamics Chemistry Study Guide Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Thermodynamics Chemistry Study Guide Answers
 - Highlighting and Note-Taking Thermodynamics Chemistry Study Guide Answers
 - Interactive Elements Thermodynamics Chemistry Study Guide Answers
8. Staying Engaged with Thermodynamics Chemistry Study Guide Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Thermodynamics Chemistry Study Guide Answers
9. Balancing eBooks and Physical Books Thermodynamics Chemistry Study Guide Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Thermodynamics Chemistry Study Guide Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Thermodynamics Chemistry Study Guide Answers
 - Setting Reading Goals Thermodynamics Chemistry Study Guide Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Thermodynamics Chemistry Study Guide Answers
 - Fact-Checking eBook Content of Thermodynamics Chemistry Study Guide Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

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