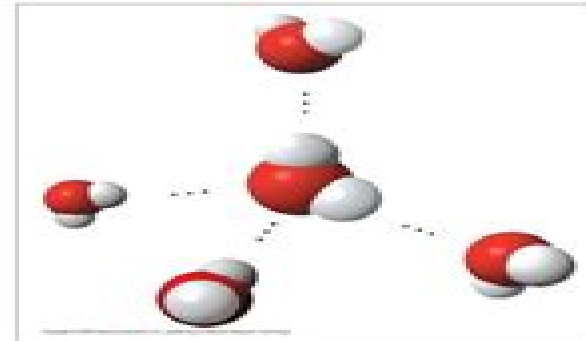


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Chapter 3: Water and Life**Concept 3.1** *The polarity of water molecules results in hydrogen bonding*

1. Study the water molecules at the right. On the central molecule, label oxygen (O) and hydrogen (H).
2. What is a *polar molecule*? Why is water considered polar?
3. Now, add + and – signs to indicate the charged regions of *each* molecule. Then, indicate the hydrogen bonds.
4. Explain *hydrogen bonding*. How many hydrogen bonds can a single water molecule form?

**Concept 3.2** *Four emergent properties of water contribute to Earth's fitness for life*

Hydrogen bonding accounts for the unique properties of water. Let's look at several.

Cohesion

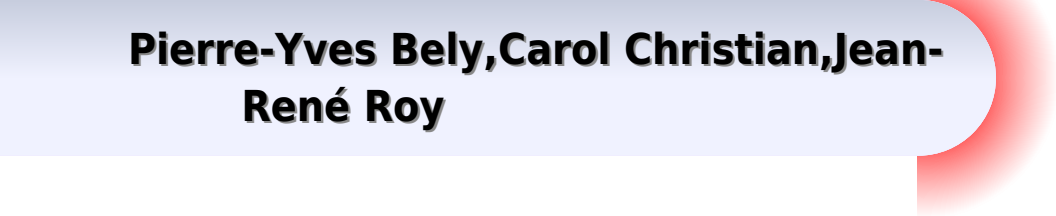
5. Distinguish between *cohesion* and *adhesion*.
6. What is demonstrated when you see beads of water on a waxed car hood?
7. Which property explains the ability of a water strider to walk on water?

Moderation of Temperature

8. The calorie is a unit of heat. Define *calorie*.
9. Water has high *specific heat*. What does this mean? How does water's specific heat compare to alcohol's?
10. Explain how hydrogen bonding contributes to water's high specific heat.

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**Pierre-Yves Bely, Carol Christian, Jean-
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Water And Life Guide Answers:

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Table of Contents Water And Life Guide Answers

1. Understanding the eBook Water And Life Guide Answers
 - The Rise of Digital Reading Water And Life Guide Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Water And Life Guide Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Water And Life Guide Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Water And Life Guide Answers

- Personalized Recommendations
 - Water And Life Guide Answers User Reviews and Ratings
 - Water And Life Guide Answers and Bestseller Lists
5. Accessing Water And Life Guide Answers Free and Paid eBooks
 - Water And Life Guide Answers Public Domain eBooks
 - Water And Life Guide Answers eBook Subscription Services
 - Water And Life Guide Answers Budget-Friendly Options
 6. Navigating Water And Life Guide Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Water And Life Guide Answers Compatibility with Devices
 - Water And Life Guide Answers Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Water And Life Guide Answers
 - Highlighting and Note-Taking Water And Life Guide Answers
 - Interactive Elements Water And Life Guide Answers
 8. Staying Engaged with Water And Life Guide Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Water And Life Guide Answers
 9. Balancing eBooks and Physical Books Water And Life Guide Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Water And Life Guide Answers
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Water And Life Guide Answers
 - Setting Reading Goals Water And Life Guide Answers
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Water And Life Guide Answers

- Fact-Checking eBook Content of Water And Life Guide Answers
- Distinguishing Credible Sources

13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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