

Chemistry 30 Thermochemistry Quiz #2 Practice

Use the following information to answer the next question:

The combustion of ethanol ($\text{C}_2\text{H}_5\text{OH}$) at 298 K, standard state ($\text{C}_2\text{H}_5\text{OH}(l)$), can be represented by the following thermochemical equation:



Equation 1

- | | |
|---|--|
| <p>8. $\text{C}_2\text{H}_5\text{OH}(g) + 3\text{O}_2(g) \rightarrow 2\text{CO}_2(g) + 3\text{H}_2\text{O}(l)$ $\Delta H^\circ = -1366.8 \text{ kJ}$</p> <p>9. $\text{C}_2\text{H}_5\text{OH}(g) + \frac{3}{2}\text{O}_2(g) \rightarrow 2\text{CO}(g) + 3\text{H}_2\text{O}(g)$ $\Delta H^\circ = -1366.8 \text{ kJ}$</p> | |
|---|--|

8. The enthalpy change for the combustion of ethanol as shown in equation 1 is:

- a. -1366.8 kJ
- b. -683.4 kJ
- c. -2733.6 kJ
- d. -2054.4 kJ

Use the following information to answer the next question:

Reaction 1 (ethanol) undergoes combustion, as represented by the following equation:



8. The maximum heat enthalpy information with the reaction is

- a. -1366.8 kJ/mol
- b. -683.4 kJ/mol
- c. -2733.6 kJ/mol
- d. -2054.4 kJ/mol

Thermochemistry Question Answer Guide

Donald J. DeCoste



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