

Thermochemistry Practice Problems ANSWER KEY

1. Calculate the standard enthalpy of reaction for the process



$$\Delta H = -176 \text{ kJ}$$

2. Calculate the heat evolved in the combustion of ethyl alcohol, ΔH_{comb}



$$\Delta H = -1366.7 \text{ kJ}$$

3. Calculate the heat required to decompose limestone, CaCO_3 , to lime, CaO , and carbon dioxide, with all substances at standard conditions.



$$\Delta H^\circ_f \text{CaCO}_3(s) = -1206.9 \text{ kJ/mol}$$

$$\Delta H^\circ_f \text{CaO}(s) = -635.1 \text{ kJ/mol}$$

$$\Delta H^\circ_f \text{CO}_2(g) = -393.5 \text{ kJ/mol}$$

$$\Delta H = 178.3 \text{ kJ}$$

4. $\text{CH}_4(g) + 3 \text{Cl}_2(g) \rightarrow \text{CHCl}_3(l) + 3 \text{HCl}(g) \quad \Delta H = ?$



$$\Delta H = -305.2 \text{ kJ}$$

5. Cyanamide, CH_2N_2 , is a weak acid that is sometimes used as a fertilizer. Calculate the standard enthalpy of formation for cyanamide, given the following standard enthalpies of reaction:



$$\Delta H = 62.1 \text{ kJ}$$

6. The standard molar enthalpy change is -905.2 kJ for the oxidation of ammonia.



Calculate the standard molar enthalpy of formation for ammonia based on the following standard enthalpies of formation: $\Delta H^\circ_f \text{NO}(g) = 90.3 \text{ kJ/mol}$ and $\Delta H^\circ_f \text{H}_2\text{O}(g) = -241.8 \text{ kJ/mol}$

$$-46.1 \text{ kJ/mol}$$

7. The thermite reaction: $2 \text{Al}(s) + \text{Fe}_2\text{O}_3(s) \rightarrow \text{Al}_2\text{O}_3(s) + 2 \text{Fe}(s) + 852 \text{ kJ}$

If you begin with 10.0 g of Al and excess Fe_2O_3 , how many kJ of heat are evolved at constant pressure?

$$158 \text{ kJ}$$

Thermochemistry Practice Calculations Answers

Mary Millhollon, Richard H. Langley



Thermochemistry Practice Calculations Answers:

Survival Guide to General Chemistry Patrick E. McMahon, Rosemary McMahon, Bohdan Khomtchouk, 2019-02-13 This work evolved over thirty combined years of teaching general chemistry to a variety of student demographics. The focus is not to recap or review the theoretical concepts well described in the available texts. Instead, the topics and descriptions in this book make available specific detailed step by step methods and procedures for solving the major types of problems in general chemistry. Explanations, instructional process sequences, solved examples, and completely solved practice problems are greatly expanded, containing significantly more detail than can usually be devoted to in a comprehensive text. Many chapters also provide alternative viewpoints as an aid to understanding. Key Features: The authors have included every major topic in the first semester of general chemistry and most major topics from the second semester. Each is written in a specific and detailed step by step process for problem solving, whether mathematical or conceptual. Each topic has greatly expanded examples and solved practice problems containing significantly more detail than found in comprehensive texts. Includes a chapter designed to eliminate confusion concerning acid base reactions, which often persists through working with acid base equilibrium. Many chapters provide alternative viewpoints as an aid to understanding. This book addresses a very real need for a large number of incoming freshmen in STEM fields.

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