

Assignment 10: Miscellaneous

1. Find the standard enthalpies of formation, ΔH_f° ,



Standard ΔH_f° for



2. A gas of bromine molecules (Br_2) is heated to 400°C and placed in a 10.0 L vessel that is initially at 1.00 bar. It is a stable equilibrium. Calculate the final temperature that is reached. $\Delta_{\text{diss}} H_f^\circ$ for Br_2 is 30.91 kJ/mol .

3. A 5.0 kg sample of aluminum initially at 20°C is submerged into 1.00 L of water at 22°C . What final temperature is reached? What is the final capacity of aluminum?

Answer: 24.88°C

4. When 1.00 g of water at 100°C is submerged into 100.0 g of water at 20°C , the final temperature is 20.88°C . What is the heat capacity of the water? What is the final capacity of water, given that the heat capacity of the surroundings is 1.00 J/K ?

5. The combustion of benzene (C_6H_6) gives heat of reaction of -3267 kJ/mol . When 1.00 g of liquid benzene is submerged into 100.0 g of water at 20°C , the final temperature is 24.88°C . Calculate the capacity of the final substance.

6. Find the standard enthalpy of formation, ΔH_f° , of the following by reaction



7. A substance heated by adding 1.00 g of water at 100°C to the sample at 20°C (initially at 20°C) reaches 24.88°C . What is the capacity of the sample? Is this process endothermic or exothermic?

8. By convention, the heat capacity of the sample is defined as the heat added to a solid body, or a solid body, divided by the temperature rise. What is the heat capacity of 1.00 g of H_2O at

Thermochemistry Practice Calculation

Joseph William Richards



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skills towards the design of their own thermochemical reactor Provides the fundamental knowledge necessary to develop fine control of thermochemical reactions **Computational Thermochemistry** Karl K. Irikura,1998 Comprises 20 contributions which grew from the August 1996 symposium Representative paper topics include estimating phase change enthalpies and entropies electrostatic covalent model parameters for molecular modeling complete basis set thermochemistry and kinetics modeling free energies of solvation and transfer use of density functional methods to compute heats of reaction and a density functional study of periodic trends in bond energies Together the contributions describe all the major methods used for estimating or predicting molecular thermochemistry Appends information on software and databases for thermochemistry essential statistical thermodynamics and worked examples Annotation copyrighted by Book News Inc Portland OR Handbook of Computational Chemistry Jerzy Leszczynski,2012-01-14 This handbook is a guide to current methods of computational chemistry explaining their limitations and advantages and providing examples of their applications The first part outlines methods the balance of volumes present numerous important applications **Chemical Calculations** Raymond Harman Ashley,1929 **Outlines of General Chemistry** Wilhelm Ostwald,1890

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