# Thermochemistry Multiple Choice Questions

|    |   | When the three types of energy changes are arranged<br>order?  A. chemical, phase, nuclear  B. nuclear, chemical, phase  |  |                           | C. nuclear, phase, chemical D. phase, chemical  |            |                                       |       |
|----|---|--|--|---------------------------|---|------------|---------------------------------------|-------|
| 0  | æ.  |  |  |                           |   |            |                                       |       |
| -  |   |  |  |                           |   |            | - 97                                  |       |
|    | 3.  | A. aluminum  | e most energy to roise I.00<br>B. copper   |                           | lead  | D.         | nickel                                |       |
| 0  | *   |  | r heat of combustion of -6<br>50 kg of water, what is the<br>8. 6.41°C   |                           |   |            | he substance is burned in<br>12.8°C   | 0.90  |
| 23 |   |  |  |                           |   |            |                                       |       |
| В  | 5.  | The addition of 9.54 kJ of heat is required to raise the temperature of 225.0 g of a liquid hydrocarbon from 20.5°C to 45.0°C. What is the heat capacity of this hydrocarbon?  |  |                           |   |            |                                       |       |
|    |   | A. 0.94 J/g-*C   | B. 1.73 7/g-*C   | C.                        | 1.88 J/g-*C   | D.         | 9.42 J/g-*C                           |       |
| c  | 6.  | If the molor heat of fusion of sodium is 2.63 kJ/mol, how much energy is needed to melt 180.0 g of solid sodium at its malting point?  |  |                           |   |            |                                       |       |
|    |   | A. 2.63 NJ   | B. 2.93 kJ   | C.                        | 20.6 kJ   | D.         | 473 kJ                                |       |
| 0  | 7.  | Which process is exertie   | rmic?  |                           |   |            |                                       |       |
|    | A. the boiling of liquid nitrogen  B. the freezing of water |  | C. The sublimetion of dry los<br>D. The vaporization of water  |                           |   |            |                                       |       |
| A  | 8   | If the heat of formation   | for the chemical reaction,   | $S_n(x) = 0$              | 2O <sub>2</sub> (g) → 85O <sub>3</sub> (g).   | e -395.7   | kJ/mel of SO <sub>3</sub> , how can 1 | that: |
|    |   | equation be rewritten to   |  |                           | a compression of the service of the service of  |            |                                       |       |
|    |   | A. S <sub>n</sub> (x) - 12O <sub>n</sub> (g) → 8SO <sub>n</sub> (g) - 3166 kJ<br>B. S <sub>n</sub> (x) - 12O <sub>n</sub> (g) → 8SO <sub>n</sub> (g) - 395.7 kJ  |  |                           | 5.(s) • 12O <sub>3</sub> (g) → 85O <sub>3</sub> (g) • 395.7 kJ 5. 5.(s) • 12O <sub>3</sub> (g) • 3166 kJ → 85O <sub>3</sub> (g) |            |                                       |       |
| 4  | 9.  | Which statement correctly describes the energy changes that take place when a solid changes to a liquid at constant temperature?  A. The potential energy increases, and the kinetic energy remains constant.  B. The potential energy remains constant, and the kinetic energy increases.  C. The potential energy decreases, and the kinetic energy decreases.  D. The potential energy increases, and the kinetic energy decreases. |  |                           |   |            |                                       |       |
| 0, | 10.   | Why does warm water have a more rapid rate of evaporation than cold water?  A. It has a higher viscosity.  B. Its molecules have a higher average kinetic energy.  C. Its molecules have more attraction for one another.  D. More of its molecules have nearly the same kinetic energy.   |  |                           |   |            |                                       |       |
| 0. | 11.   | If the heat of fusion of a substance is 20 kJ/mol, what heat is released when 1.0 mol of liquid at the melting point freezes?  |  |                           |   |            |                                       |       |
|    |   | A. 10 kJ   | B. 20 kJ   | C.                        | 40 kJ   | D.         | eo kJ                                 |       |
| C  | 12.   | initial temperature of th  | of a solid ionic compound of water was 20,0°C and the  | final tem                 | penature of the wo  | her west 4 | 7.8°C. What is Healt?                 | he    |
|    |   | A27.8 kJ/mol   | B26.2 kJ/mol   | е.                        | -13.1 kJ/mal  | (D).       | -6.26 kJ/mol                          |       |
| D  | 13.   | What does it mean if the AH value for a chemical reaction is positive?  A. kinetic energy is increasing in the system.  B. potential energy is decreasing in the system.  C. products have less potential energy than reactions.  D. reactions have less potential energy than products.   |  |                           |   |            |                                       |       |
| ^  | 14.   | Is this reaction endother  | case in the process of photo<br>6CO <sub>2</sub> (g) - 6H <sub>2</sub> O(1) - energ-<br>resic or exothermic, and is to<br>we B. endothermic near | gy it Culti<br>he value i | $\omega O_0(s) = \phi O_0(g)$<br>of $\Delta H^0$ positive or n  | egative?   | exathermic negative                   |       |

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