Apply: For a strong base, the concentration of hydroxide ions [OH] is estimated to be the same as the
concentration of the base. The pH of a strong base is found with the equation pH = 14 + log₁₀[OH]. Based
on their concentrations, find the [OH] and pH of each of the strong bases. Check your answers with the
Gizmo.

$$[NaOH] = \begin{bmatrix} 0.090 \text{ M} & [OH] = & -1.05 & pH NaOH = & 12.95 \\ [Ca(OH)_2] = & 0.021 \text{ M} & [OH]^* = & -1.68 & pH Ca(OH)_2 = & 12.32 \\ \end{bmatrix}$$

*Note: For Ca(OH)₂, the OH⁻ concentration is double the Ca(OH)₂ concentration because there are two OH⁻ ions in each Ca(OH)₂ molecule.

Activity C:	Get the Gizmo ready:	
Weak acids and bases	 Click Reset. Select 1.00 M NaOH for the Burette and Mystery CH₂COOH for the Flask. 	

Introduction: Unlike strong acids and bases, weak acids and weak bases dissociate relatively little in water. Some ions are formed, but the remaining molecules remain whole. As a result, the pH of a weak acid or base is closer to neutral than the pH of a strong acid or base.

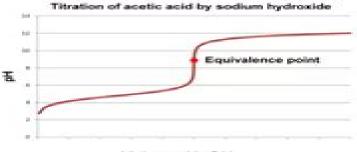
When weak acids or bases react with strong bases or acids, the resulting salts often act as bases or acids themselves, causing the pH at the equivalence point to vary from 7.0. This can impact your choice of indicator.

Question: What happens when weak acids and bases are titrated?

Gather data: For each indicator given in the Gizmo, what is the pH range over which it changes color?

Bromthymol blue: 6.0-7.6 Methyl orange: 3.1-4.4 Phenolphthalein: 8.2

 Interpret: The salt produced by the reaction of acetic acid and sodium hydroxide is a weak base, CH₃COONa. As a result, the pH of the equivalence point is slightly basic. The titration curve for this reaction is shown below:



Volume NaOH

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M Tight

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