Synthesis of Alum, KAI(SO4)2·12H2O

The purpose of this experiment is to synthesize potassium aluminam sulfate, KAl(SO₄)₂·12H₂O, commonly referred to as "alum".

Man approximately 1.8 gross of Alexandra fed. Add 27 of, KCRI. Viscours (Beir the solution and poor back time

Skeety sold 35 not. Sufferic Acid and filter again.

Allow oxystals to fixes. Filter again with Dilumel. Mass product.

Quantitative

Mass of aluminum foil	0.999 g
Mass of Sher paper (3 pieces)	3.015 g
Mass of Ep. and alum	11.194 g
Mass of alum crystals	8.179 g

Qualitative



- Write balanced net ionic equations for the following reactions which occur in this synthesis:
 - aluminum reacts with KOH and water forming potassium ions, [Al(OH)₄] and bydrogen gas

b. hydrogen ions from the acid react with the tetrahydroxosluminate ions to precipitate aluminum hydroxide

c. aluminum hydroxide reacts with additional hydrogen ions and water to form hydrated aluminum ions [Al(H₂O)_k)^{2*}

$$Al(OH)_1 + 3H^2 + 3H_2O \implies [Al(H_2O)_6]^{3}$$

d. alum forms from the potassium ions, hydrated aluminum ions, sulfate ions, and water

$$K' + [Al(H_2O)_*]^{3+} + SO_4^{2-} + 6H_2O \implies KAl(SO_4) \cdot 12H_2O$$

2. Calculate the theoretical yield of alum. What assumptions are being made in your determination of the theoretical yield?

We are assuming that the Aluminum foil is pure Aluminum and that all impurities were removed in the filtering process.

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