

CHEMISTRY Ch. 13 Notes: Water and Its Solutions

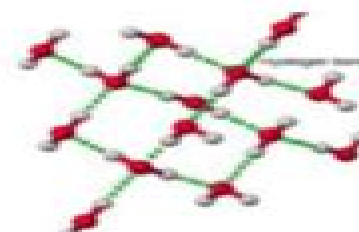
NOTE: Vocabulary terms are in **boldface and underlined**. Supporting details are in *italics*.

13.1 Notes

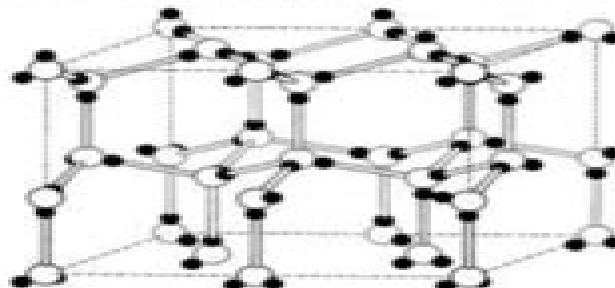
I. Water Molecule Characteristics



POLAR molecule (a dipole)



HYDROGEN BONDING in water



ICE CRYSTAL

- A) colorless and odorless
- B) neutral pH of 7
- C) *triatomic* (three atoms)
- D) *angular shape* with two unshared electron pairs
- E) *polar* (δ^+ and δ^- areas)
- F) **hydrogen bonds**—*an attraction between hydrogen and an unshared pair of an electronegative element on a neighboring molecule*
- G) **surface tension**—*attraction between molecules on the surface of a liquid*
 - 1) surface tension makes water bead
 - 2) **surfactants**—*"wetting agents" which decrease surface tension by breaking hydrogen bonds (soaps)*
- H) **capillarity (capillary action)**—*moving upward, against gravity (up through roots, etc.)*
- I) hexagonal crystals
- J) *high specific heat: 4.184 J/g°C*
- K) *high boiling point: 100 °C*

II. Heat Capacity of Water

- A) **heat capacity**—*amount of heat required to change a substance's temperature by exactly 1 °C*

Water And Aqueous Systems Chemistry Answer Key

RC Schank



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