

Worksheet 7.1 Stoichiometric Calculations

Mole to Mole Calculations

1. The octane present in gasoline burns according to the following equation.



- How many moles of O_2 are needed to react fully with 7 moles of octane (C_8H_{18})? *87.5 mol O_2*
- How many moles of CO_2 can form from 5 mole of octane? *40.0 mol CO_2*
- If the reaction is to be used to synthesize 8 moles of CO_2 , how many moles of oxygen are needed? *12.5 mol O_2*
- How many moles of carbon dioxide are formed if 17 moles of oxygen are reacted? *10.9 mol CO_2*

Mass to Mass Calculations

2. Nitrogen gas and hydrogen gas combine to form ammonia (NH_3).

- Write the balanced equation for this reaction. *$\text{N}_2 + 3\text{H}_2 \rightarrow 2\text{NH}_3$*
- If 30.0 g of nitrogen are reacted, how many moles of ammonia are produced? *2.14 mol NH_3*
- If 45.0 g of ammonia are required, what mass of hydrogen must be reacted? *7.9 g H_2*
- What mass of nitrogen is needed to react with 38.5 g of hydrogen? *179.7 g N_2*

3. Calculate how many grams of $\text{K}_2\text{Cr}_2\text{O}_7$ are needed to make 50.00 g of I_2 according to the following equation. *19.3 g $\text{K}_2\text{Cr}_2\text{O}_7$*



4. Iodine chloride, ICl , can be made by the following reaction between iodine, I_2 , potassium iodate, KIO_3 , and hydrochloric acid.



Calculate how many grams of iodine are needed to prepare 35.00 g of ICl by this reaction. *21.9 g I_2*

Solution Stoichiometry

5. Phosphoric acid can be made by boiling a solution of tetraphosphorous decaoxide in water.



- If 88.6 g of P_4O_{10} was reacted, what mass of phosphoric acid was formed? *122.3 g H_3PO_4*
- If the final volume of the acid was 500 mL, what is the molarity? *2.50 M*

6. Calculate the mass of silver that could be obtained by the reaction of a large excess of copper with 500 mL of 0.625 M AgNO_3 ? *33.7 g*

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