

UNIT 4: CHEMICAL REACTIONS STUDY GUIDE

Multiple choice:

___**C**___ 1. Chemical reactions _____.

- a. simply change the form or state of the reactants.
- b. create new atoms.
- c. always produce new substances.
- d. create more matter than what they started with.

___**B**___ 2. Which of the following is the correct skeleton equation for the combination of solid zinc with liquid bromine to produce solid zinc bromide?

- a. $\text{Zn}_{2(s)} + \text{Br}_{(l)} \rightarrow \text{ZnBr}_{2(s)}$
- b. $\text{Zn}_{(s)} + \text{Br}_{2(l)} \rightarrow \text{ZnBr}_{2(s)}$
- c. $\text{Zn}_{(s)} + \text{Br}_{2(aq)} \rightarrow \text{ZnBr}_{2(s)}$
- d. $\text{ZnBr}_{2(s)} \rightarrow \text{Zn}_{(s)} + \text{Br}_{2(l)}$

___**A**___ 3. A substance that is needed to speed up a reaction but is not a product or a reactant is known as a _____.

- a. catalyst
- b. coefficient
- c. spectator ion
- d. steroid

___**D**___ 4. In a combustion reaction, one of the reactants is always _____.

- a. a liquid
- b. water
- c. a metal
- d. oxygen

___**B**___ 5. Which of the following statements is true about the decomposition of a simple binary compound?

- a. The products cannot be predicted.
- b. The reactant is a single substance.
- c. The reactant must always be a molecular substance.
- d. The product is a single substance.

___**C**___ 6. The reaction $2\text{P}_2\text{O}_5 \rightarrow 4\text{P} + 5\text{O}_2$ is an example of which type of reaction?

- a. double replacement
- b. single replacement
- c. decomposition
- d. combustion
- e. synthesis

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Table of Contents Unit 4 Study Guide Chemistry

1. Understanding the eBook Unit 4 Study Guide Chemistry
 - The Rise of Digital Reading Unit 4 Study Guide Chemistry
 - Advantages of eBooks Over Traditional Books
2. Identifying Unit 4 Study Guide Chemistry
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Unit 4 Study Guide Chemistry
 - User-Friendly Interface
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 - Personalized Recommendations
 - Unit 4 Study Guide Chemistry User Reviews and Ratings
 - Unit 4 Study Guide Chemistry and Bestseller Lists
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 - Unit 4 Study Guide Chemistry Enhanced eBook Features
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 - Highlighting and Note-Taking Unit 4 Study Guide Chemistry
 - Interactive Elements Unit 4 Study Guide Chemistry
8. Staying Engaged with Unit 4 Study Guide Chemistry
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Unit 4 Study Guide Chemistry
9. Balancing eBooks and Physical Books Unit 4 Study Guide Chemistry
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Unit 4 Study Guide Chemistry
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
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- Setting Reading Goals Unit 4 Study Guide Chemistry
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Unit 4 Study Guide Chemistry
- Fact-Checking eBook Content of Unit 4 Study Guide Chemistry
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development

- Exploring Educational eBooks

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- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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