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mole H_2S x 18g H_2O i mole H_2O 149 4g H_2O is your theoretical yield so percent yield $\frac{137 \text{ g } \text{H}_2\text{O}}{149 \text{ g } \text{H}_2\text{O}} \times 100 = 91.77\%$ problems 6 8 are similar to problem 5

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web a write a balanced equation for the reaction $\text{Al} + \text{O}_2 \rightarrow \text{Al}_2\text{O}_3$ 4 Al + 3 $\text{O}_2 \rightarrow 2 \text{Al}_2\text{O}_3$ b determine the theoretical yield of Al_2O_3 160.0 g Al x $\frac{1 \text{ mol Al}}{27 \text{ g/mol}} = 5.93 \text{ mol}$ $\frac{160 \text{ g } \text{Al}_2\text{O}_3}{101.96 \text{ g/mol}} = 1.57 \text{ mol}$ $\frac{160 \text{ g } \text{Al}_2\text{O}_3}{302.3 \text{ g/mol}} = 0.53 \text{ mol}$ $\frac{160 \text{ g } \text{Al}_2\text{O}_3}{26.98 \text{ g/mol}} = 5.93 \text{ mol}$ 4 mol Al 1 mol Al_2O_3 c determine the percent yield $\frac{\text{actual}}{\text{theoretical}} \times 100 = \frac{260.0 \text{ g}}{302.3 \text{ g}} \times 100 = 86.01\%$

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