

**Directions:**

- You must bring a pencil and calculator to the exam.
- You will be given a periodic table.
- If you complete the entire review sheet by the day of your exam you will be eligible for the curve should one be applied.
- The answers to this review must be on a separate sheet of lined paper. You will first see the objective of each question followed by the question itself. The question is what you are providing the answer to.

**Objectives and Questions:**

**Chapter 1: Thinking Like a Scientist**

- 1) The labs you do help you process and analyze information. They also teach you how to design labs. Be aware of how lab design works. Think critically about the errors that can occur.
  - a. Why are we do need multiple trials? Constants? Controls? Variables? Hypotheses?
  - b. Define accuracy and precision. It is possible to collect data that is both accurate and precise? Is it possible to collect data that is precise but not accurate?
  - c. What can you do when you run a lab to eliminate errors? Please make sure to refer to the equipment, the variables, and the materials used for the lab.
  - d. Can theories be modified? How has the atomic theory been modified over time?

**Chapter 2: Atoms**

- 1) Know how to identify a compound as an ionic or covalent compound. Covalent compounds are also referred to as molecular compounds.
  - a. What types of elements are in binary ionic compounds? Covalent Compounds (also called binary molecular compounds)?
  - b. Circle the following formulas that represent ionic compounds. Put a box around those that are covalent/molecular.

$\text{Pd}_2\text{S}$     $\text{NaCl}$     $\text{BaCl}_2$     $\text{O}_2$     $\text{FeSO}_4$     $\text{H}_2\text{O}_2$     $\text{Fe}_2\text{O}_3$     $(\text{NH}_4)_2\text{CO}_3$     $\text{PH}_3$

  - a. Write the names and formulas for ionic compounds, covalent compounds, acids and bases.
    - i.  $\text{SnO}_2$
    - ii.  $\text{H}_2\text{S}$
    - v.  $\text{KOH}$
    - vi.  $\text{Mg}(\text{CN})_2$
    - ii.  $\text{Fe}_2(\text{PO}_4)_3$
    - iv.  $\text{HNO}_3$
    - vi.  $\text{CO}$
    - viii.  $\text{HNO}_2$
  - b. Write the formulas for the following compounds:
    - i. Silicon dioxide
    - ii. Phosphoric acid
    - iii. Lithium sulfite
    - iv. Tricarbon octahydride
    - v. potassium fluoride
    - vi. sodium hypochlorite
    - vii. Hydrocyanic acid
    - viii. Iron (II) oxide
  - c. Why do certain compounds require a roman numeral? What block of elements requires them?
- 2) Be able to predict the products of an acid and base reaction. Know what pH is and how to classify a solution as acidic, basic or neutral. Be familiar with common methods of pH detection.
  - a. Based on the following pH's, classify the solutions as acidic, basic, or neutral.
    - i.  $\text{pH} = 2$
    - ii.  $\text{pH} = 5.8$
    - iii.  $\text{pH} = 7$
  - b. List two properties each for acids and bases. (Examples include taste, what ion they create, feel, and ability to react with a metal.)
  - c. What is more acidic, a solution with a pH of 2 or 5? Which is more basic, a pH of 8 or 14?

# Utica Community Schools Chemistry Final Review

**PT Brinkman**



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## **Table of Contents Utica Community Schools Chemistry Final Review**

1. Understanding the eBook Utica Community Schools Chemistry Final Review
  - The Rise of Digital Reading Utica Community Schools Chemistry Final Review
  - Advantages of eBooks Over Traditional Books
2. Identifying Utica Community Schools Chemistry Final Review
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Utica Community Schools Chemistry Final Review
  - User-Friendly Interface
4. Exploring eBook Recommendations from Utica Community Schools Chemistry Final Review
  - Personalized Recommendations
  - Utica Community Schools Chemistry Final Review User Reviews and Ratings

- Utica Community Schools Chemistry Final Review and Bestseller Lists
- 5. Accessing Utica Community Schools Chemistry Final Review Free and Paid eBooks
  - Utica Community Schools Chemistry Final Review Public Domain eBooks
  - Utica Community Schools Chemistry Final Review eBook Subscription Services
  - Utica Community Schools Chemistry Final Review Budget-Friendly Options
- 6. Navigating Utica Community Schools Chemistry Final Review eBook Formats
  - ePub, PDF, MOBI, and More
  - Utica Community Schools Chemistry Final Review Compatibility with Devices
  - Utica Community Schools Chemistry Final Review Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Utica Community Schools Chemistry Final Review
  - Highlighting and Note-Taking Utica Community Schools Chemistry Final Review
  - Interactive Elements Utica Community Schools Chemistry Final Review
- 8. Staying Engaged with Utica Community Schools Chemistry Final Review
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Utica Community Schools Chemistry Final Review
- 9. Balancing eBooks and Physical Books Utica Community Schools Chemistry Final Review
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Utica Community Schools Chemistry Final Review
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Utica Community Schools Chemistry Final Review
  - Setting Reading Goals Utica Community Schools Chemistry Final Review
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Utica Community Schools Chemistry Final Review
  - Fact-Checking eBook Content of Utica Community Schools Chemistry Final Review
  - Distinguishing Credible Sources

13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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web apr 17 2023 the logiks general intermediate test is a timed test there are 50 questions in total the test is split into three parts with each part assessing a specific ability numerical verbal reasoning and abstract reasoning