

When You Comin Back,  
Red Ryder.®

# When You Comin Back Red Ryder

**Mark Howard Medoff**



## **When You Comin Back Red Ryder:**

**When You Coming Back, Red Ryder** Kenneth Frankel, Elliot Martin, Brad Dourif, E. J. Oshins, Elizabeth Sturges, Kevin Conway, Addison Powell, Joe Jamrog, Robyn Goodman, James Kiernan, Kristin Van Buren, Circle Repertory Company, Eastside Playhouse (Organization : New York, N.Y.), 1973      **When You Comin' Back, Red Ryder?** Mark Medoff, Theatre Passe Muraille Archives, 1979      **When You Comin' Back, Red Ryder?** Mark Howard Medoff, 1974-10 THE STORY The scene is an all night diner in a sleepy southwestern town the time early Sunday morning when the night attendant young Stephen Red Ryder is about to turn his duties over to his daytime counterpart Angel Her friend Lyle who r      **When You Comin Back, Red Ryder?** Mark Howard Medoff, 1974 A taut and gripping story of confrontation in its rawest aspects Into a run down diner in a small New Mexico border town is intruded a mysterious and ominous figure Teddy with his girl Cheryl He has run out of funds and his VW van needs repair In the diner he finds Angel and Stephen Red who work there the cripple Lyle and Richard and Clarisse the successful businessman and his violinist wife Teddy terrorizes this group the portrayal of his machinations provides on the surface a suspense story of great impact Dust jacket      **When You Comin Back, Red Ryder?** Mark Howard Medoff, 1974      **When You Comin Back, Red Ryder?** Mark Howard Medoff, 1974      *A Production of Mark Medoff's When You Comin' Back, Red Ryder?* Dwight E. Watson, 1979      **A Production of Mark Medoff's When You Comin Back, Red Ryder?** Susan Marie Barrett, 1981      *When You Comin' Back, Red Ryder?* Carole Jeanne Palmer, 1989  
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web donor whereas a base generically represented by  $b$  is a proton acceptor when acids and bases combine they swap a proton  $h_a b + h_b$  strong acids e.g. sulfuric acid have a very pronounced tendency to transfer a proton to a base whereas weak acids

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web understand that parts of an acid base titration be able to determine the  $K_a$  or  $K_b$  from pH data associated with the titration of a weak acid or base be able to determine the molar mass of a solid monoprotic acid from titration data be able to calculate  $K_{a1}$  and  $K_{a2}$  for a polyprotic acid by the end of this lab students should be able to

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web 1 approximately 3 g of hydrated oxalic acid  $\text{H}_2\text{C}_2\text{O}_4 \cdot 2\text{H}_2\text{O}$  was weighted to the nearest 0.01 g in a 50 ml beaker 2 about 30 ml of distilled water was added to the beaker to dissolve the oxalic acid 3 the solution was transferred into 250 ml volumetric flask the beaker was rinsed and its contents were poured into the flask

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web solution a titrant volume 0 ml the solution ph is due to the acid ionization of hcl because this is a strong acid the ionization is complete and the hydronium ion molarity is 0.100 M the ph of the solution is then  $\text{pH} = -\log 0.100 = 1.000$  ph log 0.100 1.000 b titrant volume 12.50 ml

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