

3.2.10 aromatic chemistry

structure of benzene

- benzene formula: C_6H_6
- planar cyclic molecule and non-polar
- each carbon is bonded to two other carbons in a ring and one hydrogen
- the bond lengths between the carbon atoms in benzene are equal, and hence the bonds are identical
- π -electrons are made up of the carbon atom orbitals to form a ring of delocalised electron (delocalised system) or resonance
- all the bonds in the ring are the same length. $\rightarrow$ trigonal planar which is an indication of the $C-C$ (134pm) and $C-H$ (109pm)



ring of delocalised electrons can be represented by a circle in the ring of carbon-carbon bonds (circle model)

delocalised π system means benzene more stable



stability of benzene

benzene's unusual structure: Chemical systems $\rightarrow$ 1,3,5-triene



benzene proposed that benzene occurred in a ring and single bonds

problem with this structure:

- benzene doesn't undergo addition reactions
- benzene is "aromatic" more stable than triene (benzene's unusual structure $\rightarrow$ 1,3,5-triene)
- bond lengths: $C-C$ (134pm) $C-H$ (109pm) in benzene all bonds are 139pm

what unexpected structural change of hydrogenation for benzene with the normal change of hydrogenation for triene? but the heat released is not even similar



hydrogenation for the double bond
this is $\rightarrow$ hydrogenated the structural change is $\rightarrow$ "aromatic"



if benzene had three double bonds
expected structural change of hydrogenation would be $\rightarrow$ "aromatic"



benzene the experimental structural change of hydrogenation of benzene is $\rightarrow$ "aromatic"

benzene is more stable than 1,3,5-triene
the heat released is not even similar

aromaticity: definition

benzene: aromatic molecule in benzene ring

what happens if the whole structural group is
put in a structural group: stable $\rightarrow$ benzene



what happens if you put the whole structural group in
a group: stable $\rightarrow$ benzene



benzene is aromatic

the structure of benzene is not a Kekulé structure
it is a resonance structure
it is a structure that is not a Kekulé structure

aromaticity

benzene is not aromatic

benzene is aromatic because it has a delocalised π system $\rightarrow$ extra stability
 $\rightarrow$ 139pm $\rightarrow$ 134pm

benzene has the same energy as benzene structure $\rightarrow$ it is not aromatic

it is not aromatic because

experimental structural change of benzene is "aromatic" and structure that is not

energy is required to break bonds $\rightarrow$ more energy is needed to break the double bond in benzene than in the substituted alkenes $\rightarrow$ 1,3,5-triene

benzene is very stable in it has not undergone hydrogenation addition reactions as it is aromatic the ring of delocalised electrons

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 - The Rise of Digital Reading Wjec Hb4 January 2011 June Mark Schemes
 - Advantages of eBooks Over Traditional Books
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 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
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 - Features to Look for in an Wjec Hb4 January 2011 June Mark Schemes
 - User-Friendly Interface
4. Exploring eBook Recommendations from Wjec Hb4 January 2011 June Mark Schemes
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 - Wjec Hb4 January 2011 June Mark Schemes and Bestseller Lists
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 - Wjec Hb4 January 2011 June Mark Schemes eBook Subscription Services
 - Wjec Hb4 January 2011 June Mark Schemes Budget-Friendly Options
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- Wjec Hb4 January 2011 June Mark Schemes Compatibility with Devices
- Wjec Hb4 January 2011 June Mark Schemes Enhanced eBook Features
- 7. Enhancing Your Reading Experience
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 - Highlighting and Note-Taking Wjec Hb4 January 2011 June Mark Schemes
 - Interactive Elements Wjec Hb4 January 2011 June Mark Schemes
- 8. Staying Engaged with Wjec Hb4 January 2011 June Mark Schemes
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Wjec Hb4 January 2011 June Mark Schemes
- 9. Balancing eBooks and Physical Books Wjec Hb4 January 2011 June Mark Schemes
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Wjec Hb4 January 2011 June Mark Schemes
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
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 - Setting Reading Goals Wjec Hb4 January 2011 June Mark Schemes
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Wjec Hb4 January 2011 June Mark Schemes
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 - Distinguishing Credible Sources
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 - Exploring Educational eBooks
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 - Integration of Multimedia Elements
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